

# Acids & Bases HOMEWORK

## Homework

1. Propose the correct ion search of the strong acid listed below. Propose the incorrect ion search of the strong base listed below.

- a. **HF**: **Hydrofluoric acid**
- b. **HI**: **Hydroiodic acid**
- c. **HNO<sub>3</sub>**: **nitric acid**
- d. **HClO<sub>4</sub>**: **perchloric acid**
- e. **HBrO<sub>3</sub>**: **hypobromous acid**
- f. **HClO<sub>4</sub>**: **chloric acid**
- g. **H<sub>2</sub>SO<sub>4</sub>**: **hydrochloric acid**
- h. **H<sub>2</sub>SO<sub>4</sub>**: **hydrobromic acid**
- i. **H<sub>2</sub>SO<sub>4</sub>**: **hydroiodic acid**
- j. **H<sub>2</sub>SO<sub>4</sub>**: **hydrochloric acid**

2. Of the acids and bases listed below, write the ones that completely dissociate in water.

Acid	Base	Mixture	Neutral	Acidic	Basic	Neutral	Basic
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3. Write the dissociation for **Na<sub>2</sub>CO<sub>3</sub>**.

4. Propose both ion search terms when asked to find **HF** in the periodic table.

- a. Write a balanced equation for this reaction:  $\text{H}_2\text{O} + \text{HF} \longrightarrow \text{H}_3\text{O}^+ + \text{F}^-$

5. Propose both ion search terms when asked to find **NaOH** in the periodic table.

6. Propose both ion search terms when asked to find **NaClO<sub>4</sub>** in the periodic table.

7. In your ion search, is **NaClO<sub>4</sub>** acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid**. Because it is a proton acceptor.

8. Identify the conjugate base of **NaClO<sub>4</sub>**: **ClO<sub>4</sub><sup>-</sup>**

9. Identify **NaClO<sub>4</sub>** as a protonator, deprotonator, or neutral. **Protonator**. Because it is a strong base. No deprotonation.

10. **HF** will react with water and form a **HF<sub>2</sub><sup>-</sup>** ion in the process.

- a. Write a balanced equation for this reaction:  $\text{H}_2\text{O} + \text{HF} \longrightarrow \text{H}_3\text{O}^+ + \text{HF}_2^-$

11. Propose both ion search terms when asked to find **Na<sub>2</sub>CO<sub>3</sub>** in the periodic table.

12. Propose both ion search terms when asked to find **NaOH** in the periodic table.

13. In your ion search, is **NaOH** acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base**. Because it is a proton acceptor.

14. Identify the conjugate acid of **HF**: **HF<sub>2</sub><sup>-</sup>**

15. Identify **HF** as a protonator, deprotonator, or neutral. **Protonator**

16. **HF<sub>2</sub><sup>-</sup>** will react with water and form a **HF<sub>2</sub><sup>-</sup>** ion in the process.

- a. Write a balanced equation for this reaction:  $\text{H}_2\text{O} + \text{HF}_2^- \longrightarrow \text{H}_3\text{O}^+ + \text{HF}_2^-$

17. Propose both ion search terms when asked to find **Na<sub>2</sub>CO<sub>3</sub>** in the periodic table.

18. Propose both ion search terms when asked to find **NaOH** in the periodic table.

19. In your ion search, is **NaOH** acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base**. Because it is a proton acceptor.

20. Identify the conjugate base of **HF<sub>2</sub><sup>-</sup>**: **HF<sub>2</sub><sup>-</sup>**

21. Identify **HF<sub>2</sub><sup>-</sup>** as a protonator, deprotonator, or neutral. **Protonator**. Because it is a strong base. No deprotonation.

# 19 3 Acids And Bases Answer Key Chemistry

**James Vincent Quagliano, L. M.  
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