

TEST YOURSELF

TEST YOURSELF 2.1 (THEORIES AND EXPERIMENTS RELATED TO STRUCTURE OF ATOMS)

i) Do you know any element having no neutrons in its atoms?

Ans: Yes the isotope of hydrogen (Protium, ${}^1\text{H}$) has no neutron. It has one proton in its nucleus and one electron revolving around it.

ii) Who discovered an electron, a proton and a neutron?

Ans: a. In 1886, Goldstein discovered positively charged particles called protons.
b. In 1897, J.J. Thomson discovered the negatively charged particles called electrons.
c. In 1932, James Chadwick discovered the neutral particles called neutrons.

iii) How does electron differ from a neutron?

Ans: Difference between electron and neutron.

Electron	Neutron
i. Electron is the negatively charged particle.	i. Neutron is the neutral particle.
ii. Electron revolves around the nucleus.	ii. Neutron is present in the nucleus.
iii. Mass of electron is 5.486×10^{-5} g amu or 9.106×10^{-28} g.	iii. Mass of neutron is 1.0087 amu or 1.674×10^{-24} g.

iv) Explain how anode rays are formed from the gas taken in the discharge tube?

Ans: Anode rays do not originate from the anode. In fact these rays are produced when the cathode rays or electrons collide with the residual gas molecules present in the discharge tube and ionize them as follows:



TEST YOURSELF 2.2 RUTHERFORD'S ATOMIC MODEL AND BOHR'S ATOMIC MODEL

i) How was it proved that the whole mass of an atom is located at its centre?

Ans: Rutherford bombarded alpha particles on a 0.00004 cm thick gold foil. Almost all the particles passed through the foil undeflected, only few bounced back. The complete rebounce of few alpha particles showed that the nucleus is an extremely small positively charged part. It is situated at the center of an atom and it carries nearly the whole mass of an atom.

ii) How was it shown that atomic nuclei are positively charged?

Ans: Alpha particles are helium nuclei (He^{2+}) i.e. doubly positively charged particles. The deflection and rebounding of a few particles showed that atomic nuclei (centre of an atom) are positively charged.

iii) Name the particles which determine the mass of an atom.

Ans: Protons and neutrons present in the nucleus determine the mass of an atom. The sum of number of protons and neutrons is equal to mass of an atom.

iv) What is the classical theory of radiation? How does it differ from quantum theory?

Ans:

Classical theory of radiation	Quantum theory of radiation
According to classical theory of radiation electrons being the charged particles should release or emit energy continuously and they should ultimately fall into the nucleus.	According to quantum theory of radiation, energy is not emitted or absorbed continuously but it is emitted or absorbed in the form of small energy packets or bundles known as quantum or

Quiz Yourself Chapter 8 Introduction To Chemistry

**Kevin T. Patton, Frank B. Bell, Terry
Thompson, Peggy L. Williamson**

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Fachenglisch für Laborberufe

Steven L. Hanft, 2015-09-15 Mit fortschreitender Globalisierung von Waren und Dienstleistungen hat es an immer mehr Arbeitsplätzen in Chemie, Pharma und Biotech Branche die englische Sprache Einzug. In der Schule hat man zwar gelernt sich über Alltagsthemen zu unterhalten, aber wenn es darum geht dem Kundendienst am Telefon die Fehlfunktion des teuersten Geräts im Labor zu beschreiben, kommt doch so mancher ins Schwitzen. Nach einer Einführung in der die wichtigsten Besonderheiten der englischen Sprache aus Sicht eines deutschen Sprechers rekapituliert werden, behandelt der Autor in 14 Lektionen Schritt für Schritt den Spezialwortschatz und fachspezifische Sprach- und Schreibformen. Die Themen reichen von mathematischen Ausdrücken über chemische Nomenklatur, Biomoleküle, Versuchstiere und Prozesstechnik bis hin zum Umgang mit Regulierungsbehörden und Audits. Gesprächssituationen wie der Anruf beim Kundendienst, die Vorstellung beim neuen Chef oder das Kundengespräch am Messestand werden analysiert und eingebettet. Mit direktem Bezug zur Berufspraxis geht dieser Sprachführer über herkömmliche Englischkurse weit hinaus und bietet wertvolle Hilfe für alle, die im Beruf besser Englisch sprechen wollen. Auch für den fachbezogenen Sprachunterricht an Fachschulen und Hochschulen ist dieses Buch bestens geeignet. Komplett mit Bürgern, Tests und Rezepten, wie man die häufigsten Fehler vermeidet. Das Buch ist auch als e-Book mit Audiounterstützung erhältlich.

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Quiz Refresher in Chemistry Rolland J. Gladieux,1953

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