

TEST YOUR SELF

TEST YOURSELF 2.1 (THEORIES AND EXPERIMENTS RELATED TO STRUCTURE OF ATOMS)

i) Do you know any element having no neutrons in its atoms?

Ans: Yes the isotope of hydrogen (Protium, ${}^1_1\text{H}$) has no neutron. It has one proton in its nucleus and one electron revolving around it.

ii) Who discovered an electron, a proton and a neutron?

Ans: a. In 1886, Goldstein discovered positively charged particles called protons.

b. In 1897, J.J. Thomson discovered the negatively charged particles called electrons.

c. In 1932, James Chadwick discovered the neutral particles called neutrons.

iii) How does electron differ from a neutron?

Ans: Difference between electron and neutron.

Electron	Neutron
i. Electron is the negatively charged particle.	i. Neutron is the neutral particle.
ii. Electron revolves around the nucleus.	ii. Neutron is present in the nucleus.
iii. Mass of electron is 5.486×10^{-4} g or 9.106×10^{-28} g.	iii. Mass of neutron is 1.0087 amu or 1.674×10^{-24} g.

iv) Explain how anode rays are formed from the gas taken in the discharge tube?

Ans: Anode rays do not originate from the anode. In fact these rays are produced when the cathode rays or electrons collide with the residual gas molecules present in the discharge tube and ionize them as follows:



TEST YOURSELF 2.2 RUTHERFORD'S ATOMIC MODEL AND BOHR'S ATOMIC MODEL

i) How was it proved that the whole mass of an atom is located at its centre?

Ans: Rutherford bombarded alpha particles on a 0.00004 cm thick gold foil. Almost all the particles passed through the foil undeflected, only few bounced back. The complete rebound of few alpha particles showed that the nucleus is an extremely small positively charged part. It is situated at the center of an atom and it carries nearly the whole mass of an atom.

ii) How was it shown that atomic nuclei are positively charged?

Ans: Alpha particles are helium nuclei (He^{2+}) i.e. doubly positively charged particles. The deflection and rebounding of a few particles showed that atomic nuclei (centre of an atom) are positively charged.

iii) Name the particles which determine the mass of an atom.

Ans: Protons and neutrons present in the nucleus determine the mass of an atom. The sum of number of protons and neutrons is equal to mass of an atom.

iv) What is the classical theory of radiation? How does it differ from quantum theory?

Ans:

Classical theory of radiation	Quantum theory of radiation
According to classical theory of radiation electrons being the charged particles should release or emit energy continuously and they should ultimately fall into the nucleus.	According to quantum theory of radiation, energy is not emitted or absorbed continuously but it is emitted or absorbed in the form of small energy packets or bundles known as quantum or

Quiz Yourself Chapter 8 Introductcion To Chemistry

**Kevin T. Patton, Frank B. Bell, Terry
Thompson, Peggie L. Williamson**

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Dickerson, Gray, Darensbourg, and Darensbourg Patricia L. Samuel, 1984 Timetable University of Illinois at Urbana-Champaign, 2000 **Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e** Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond s text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry English Mechanic and Mirror of Science and Art ,1883 **Complete Preparation for the MCAT** Williams & Wilkins Review, 1998-04 Here is the most respected test prep book for the Medical College Admission Test you can buy featuring an active learning approach for a better understanding of the exam s content and a better chance for success Unique to this guide are coverage of all recent changes in the MCAT plus a step by step plan for sharpening cognitive skills developing problem solving skills and critical thinking This thorough guide replaces expensive test preparation courses while giving students exactly what they need to get ready for the MCAT **Rhetoric for Academic Reasoning** L. Bense-Meyers, 1992 **Teaching Science for All Children** ,2005 This compact paperback volume provides preservice teachers with STRATEGIES AND METHODS of teaching science in the K 8 classroom using Inquiry The authors integrate the NSE standards constructivism and technology into their popular E approach to teaching Exploration Explanation Expansion and Evaluation make up the 4 E s of the learning cycle model first invented by Robert Karplus as part of the Science Curriculum Improvement Study in the 1960s Teaching Science for All Children Inquiry Methods for Constructing Understanding provides methods for future teachers to foster awareness among their students of the nature of science to implement skills in the classroom using science inquiry processes and to develop in their students an understanding of the interactions among science technology and society English Mechanic and World of Science ,1879 Report[s], [minutes of Evidence, Indexes, Answers to Questions]. Great Britain. Royal Commission on Labour, 1894 **The Journal of Education** ,1899 *A Complete Preparation for the MCAT* Aftab S. Hassan, James L. Flowers, 1992 This guide for MCAT preparation applies the principles of active and problem based learning to an updated review of content and skills with models for enhanced problem solving and critical thinking abilities There are details on setting up a self managed study programme with guidelines for time management and stress management All areas tested on the exam are covered verbal reasoning physical science writing sample biological sciences with practice questions to chart progress

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