

Multiple Choice:

Selecting the alternative that constitutes the statement or answers the question. What is most commonly observed with water?

1. How many moles of hydrogen gas are in a cylinder at typical conditions, 1000 L?
 a. 1
 b. 10
 c. 100
 d. 1000

2. At temperature $T_1 = 100^\circ\text{C}$, water has a vapor pressure of 1000 mmHg.
 a. 100
 b. 1000
 c. 10000
 d. 100000

3. The condensed phase is in the lowest energy state.
 a. Most of the time
 b. Only at low temperatures
 c. Only at high temperatures
 d. Only at intermediate temperatures

4. For temperature $T_1 = 100^\circ\text{C}$, $\Delta H_{\text{vap}} = 40.7 \text{ kJ/mol}$, how many moles of water vapor are produced when 1000 J of heat is added?
 a. 100 moles
 b. 1000 moles
 c. 10000 moles
 d. 100000 moles

5. What is the state of the water molecules in the condensed phase?
 a. Disordered
 b. Ordered
 c. Intermediate
 d. None of the above

6. What is the boiling point of water at 1000 mmHg?
 a. 100°C
 b. 100°C
 c. 100°C
 d. 100°C

7. A sample of water is heated to 100°C , at this point, the water is at the boiling point. What is the state of the water?
 a. Solid
 b. Liquid
 c. Gas
 d. None of the above

8. For temperature $T_1 = 100^\circ\text{C}$, $\Delta H_{\text{vap}} = 40.7 \text{ kJ/mol}$, how many moles of water vapor are produced when 1000 J of heat is added?
 a. 100 moles
 b. 1000 moles
 c. 10000 moles
 d. 100000 moles

9. At temperature $T_1 = 100^\circ\text{C}$, $\Delta H_{\text{vap}} = 40.7 \text{ kJ/mol}$, how many moles of water vapor are produced when 1000 J of heat is added?
 a. 100 moles
 b. 1000 moles
 c. 10000 moles
 d. 100000 moles

10. What is the boiling point of water at 1000 mmHg?
 a. 100°C
 b. 100°C
 c. 100°C
 d. 100°C

11. The boiling point of water is 100°C at 1000 mmHg. What is the boiling point of water at 1000 mmHg?

Multiple Choice Stoichiometry Test

R Bogdan



Multiple Choice Stoichiometry Test:

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