

## Practicing the Mole -- Even Problems Answer Key

*Calculate the mass in grams of each of the following:*

2. 8.00 moles of aluminum

$$\frac{8.00 \text{ mol}}{1 \text{ mol}} \left| \begin{array}{c} 26.98 \text{ g} \\ \hline \end{array} \right| = 216 \text{ g Al}$$

6. 7.00 moles of iodine ( $I_2$ )

$$\frac{7.00 \text{ mol}}{1 \text{ mol}} \left| \begin{array}{c} 253.8 \text{ g} \\ \hline \end{array} \right| = 1,780 \text{ g } I_2$$

4.  $2.00 \times 10^2$  moles of chlorine ( $Cl_2$ )

$$\frac{200. \text{ mol}}{1 \text{ mol}} \left| \begin{array}{c} 70.90 \text{ g} \\ \hline \end{array} \right| = 14,200 \text{ g } Cl_2$$

8. 9.20 moles of iron

$$\frac{9.20 \text{ mol}}{1 \text{ mol}} \left| \begin{array}{c} 55.85 \text{ g} \\ \hline \end{array} \right| = 514 \text{ g Fe}$$

*Calculate the number of moles of atoms in each of the following:*

10. 281 g of iron

$$\frac{281 \text{ g}}{55.85 \text{ g}} \left| \begin{array}{c} 1 \text{ mol} \\ \hline \end{array} \right| = 5.03 \text{ mol Fe}$$

14. 93.0 g of phosphorus

$$\frac{93.0 \text{ g}}{30.97 \text{ g}} \left| \begin{array}{c} 1 \text{ mol} \\ \hline \end{array} \right| = 3.00 \text{ mol P}$$

12. 72.0 g of magnesium

$$\frac{72.0 \text{ g}}{24.31 \text{ g}} \left| \begin{array}{c} 1 \text{ mol} \\ \hline \end{array} \right| = 2.96 \text{ mol Mg}$$

16. 1035 g of lead

$$\frac{1035 \text{ g}}{207.2 \text{ g}} \left| \begin{array}{c} 1 \text{ mol} \\ \hline \end{array} \right| = 4.995 \text{ mol Pb}$$

*Calculate the number of atoms in each of the following:*

18. 5.40 moles of iron

$$\frac{5.40 \text{ mol}}{1 \text{ mole}} \left| \begin{array}{c} 6.022 \times 10^{23} \text{ atoms} \\ \hline \end{array} \right|$$

$$= 3.25 \times 10^{24} \text{ atoms Fe}$$

20. 10.0 moles of titanium

$$\frac{10.0 \text{ mol}}{1 \text{ mole}} \left| \begin{array}{c} 6.022 \times 10^{23} \text{ atoms} \\ \hline \end{array} \right|$$

$$= 6.02 \times 10^{24} \text{ atoms Ti}$$

# The Practice Of Chemistry

**Wink,Sharon Fetzer-Gislason,McNichol**

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## Table of Contents The Practice Of Chemistry

1. Understanding the eBook The Practice Of Chemistry
  - The Rise of Digital Reading The Practice Of Chemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying The Practice Of Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an The Practice Of Chemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from The Practice Of Chemistry
  - Personalized Recommendations
  - The Practice Of Chemistry User Reviews and Ratings
  - The Practice Of Chemistry and Bestseller Lists
5. Accessing The Practice Of Chemistry Free and Paid eBooks
  - The Practice Of Chemistry Public Domain eBooks
  - The Practice Of Chemistry eBook Subscription Services
  - The Practice Of Chemistry Budget-Friendly Options
6. Navigating The Practice Of Chemistry eBook Formats

- ePub, PDF, MOBI, and More
- The Practice Of Chemistry Compatibility with Devices
- The Practice Of Chemistry Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of The Practice Of Chemistry
- Highlighting and Note-Taking The Practice Of Chemistry
- Interactive Elements The Practice Of Chemistry

8. Staying Engaged with The Practice Of Chemistry

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers The Practice Of Chemistry

9. Balancing eBooks and Physical Books The Practice Of Chemistry

- Benefits of a Digital Library
- Creating a Diverse Reading Collection The Practice Of Chemistry

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine The Practice Of Chemistry

- Setting Reading Goals The Practice Of Chemistry
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of The Practice Of Chemistry

- Fact-Checking eBook Content of The Practice Of Chemistry
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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descriptive measures 6 probability theory and probability distributions 7 sampling distributions 8 inferential statistics 9

hypothesis testing 10 inference from small samples 11 analysis of variance anova

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