

TEST YOUR SELF

TEST YOURSELF 2.1 (THEORIES AND EXPERIMENTS RELATED TO STRUCTURE OF ATOMS)

i) Do you know any element having no neutrons in its atoms?

Ans: Yes the isotope of hydrogen (Protium, ${}^1_1\text{H}$) has no neutron. It has one proton in its nucleus and one electron revolving around it.

ii) Who discovered an electron, a proton and a neutron?

Ans: a. In 1886, Goldstein discovered positively charged particles called protons.

b. In 1897, J.J. Thomson discovered the negatively charged particles called electrons.

c. In 1932, James Chadwick discovered the neutral particles called neutrons.

iii) How does electron differ from a neutron?

Ans: Difference between electron and neutron.

Electron	Neutron
i. Electron is the negatively charged particle.	i. Neutron is the neutral particle.
ii. Electron revolves around the nucleus.	ii. Neutron is present in the nucleus.
iii. Mass of electron is 5.486×10^{-4} g or 9.106×10^{-28} g.	iii. Mass of neutron is 1.0087 amu or 1.674×10^{-24} g.

iv) Explain how anode rays are formed from the gas taken in the discharge tube?

Ans: Anode rays do not originate from the anode. In fact these rays are produced when the cathode rays or electrons collide with the residual gas molecules present in the discharge tube and ionize them as follows:



TEST YOURSELF 2.2 RUTHERFORD'S ATOMIC MODEL AND BOHR'S ATOMIC MODEL

i) How was it proved that the whole mass of an atom is located at its centre?

Ans: Rutherford bombarded alpha particles on a 0.00004 cm thick gold foil. Almost all the particles passed through the foil undeflected, only few bounced back. The complete rebound of few alpha particles showed that the nucleus is an extremely small positively charged part. It is situated at the center of an atom and it carries nearly the whole mass of an atom.

ii) How was it shown that atomic nuclei are positively charged?

Ans: Alpha particles are helium nuclei (He^{2+}) i.e. doubly positively charged particles. The deflection and rebounding of a few particles showed that atomic nuclei (centre of an atom) are positively charged.

iii) Name the particles which determine the mass of an atom.

Ans: Protons and neutrons present in the nucleus determine the mass of an atom. The sum of number of protons and neutrons is equal to mass of an atom.

iv) What is the classical theory of radiation? How does it differ from quantum theory?

Ans:

Classical theory of radiation	Quantum theory of radiation
According to classical theory of radiation electrons being the charged particles should release or emit energy continuously and they should ultimately fall into the nucleus.	According to quantum theory of radiation, energy is not emitted or absorbed continuously but it is emitted or absorbed in the form of small energy packets or bundles known as quantum or

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Lois Weinstein



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Guide to Mining Examinations with Answers to Examination Questions Henry Davies, 1894 Quiz Refresher in Chemistry Rolland J. Gladieux, 1953 Study Guide for Chemical Principles, Fourth Edition, by Dickerson, Gray, Darensbourg, and Darensbourg Patricia L. Samuel, 1984 *Timetable* University of Illinois at Urbana-Champaign, 2000

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Table of Contents Quiz Yourself Chapter 8 Introductcion To Chemistry

1. Understanding the eBook Quiz Yourself Chapter 8 Introductcion To Chemistry
 - The Rise of Digital Reading Quiz Yourself Chapter 8 Introductcion To Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Quiz Yourself Chapter 8 Introductcion To Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Quiz Yourself Chapter 8 Introductcion To Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Quiz Yourself Chapter 8 Introductcion To Chemistry
 - Personalized Recommendations
 - Quiz Yourself Chapter 8 Introductcion To Chemistry User Reviews and Ratings
 - Quiz Yourself Chapter 8 Introductcion To Chemistry and Bestseller Lists
5. Accessing Quiz Yourself Chapter 8 Introductcion To Chemistry Free and Paid eBooks
 - Quiz Yourself Chapter 8 Introductcion To Chemistry Public Domain eBooks
 - Quiz Yourself Chapter 8 Introductcion To Chemistry eBook Subscription Services
 - Quiz Yourself Chapter 8 Introductcion To Chemistry Budget-Friendly Options
6. Navigating Quiz Yourself Chapter 8 Introductcion To Chemistry eBook Formats

- ePub, PDF, MOBI, and More
- Quiz Yourself Chapter 8 Introduction To Chemistry Compatibility with Devices
- Quiz Yourself Chapter 8 Introduction To Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Quiz Yourself Chapter 8 Introduction To Chemistry
 - Highlighting and Note-Taking Quiz Yourself Chapter 8 Introduction To Chemistry
 - Interactive Elements Quiz Yourself Chapter 8 Introduction To Chemistry
- 8. Staying Engaged with Quiz Yourself Chapter 8 Introduction To Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Quiz Yourself Chapter 8 Introduction To Chemistry
- 9. Balancing eBooks and Physical Books Quiz Yourself Chapter 8 Introduction To Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Quiz Yourself Chapter 8 Introduction To Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Quiz Yourself Chapter 8 Introduction To Chemistry
 - Setting Reading Goals Quiz Yourself Chapter 8 Introduction To Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Quiz Yourself Chapter 8 Introduction To Chemistry
 - Fact-Checking eBook Content of Quiz Yourself Chapter 8 Introduction To Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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