

19.5

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and pOH and concentrations of acids and bases and vice versa.
3. Calculate the pH of weak acids.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid-base indicator |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the requested unit or numerical information of the values indicated. All answers should be given with the correct number of significant figures and the appropriate unit.

1. What substance is _____, a weak hydrogen ion donor?
2. _____ is the concentration of H^+ ions present in 0.00100 mol/L of _____.
3. _____ is the concentration of _____ in 0.00100 mol/L of _____.
4. _____ is the concentration of _____ in 0.00100 mol/L of _____.
5. _____ is the concentration of _____ in 0.00100 mol/L of _____.
6. _____ is the concentration of _____ in 0.00100 mol/L of _____.
7. _____ is the concentration of _____ in 0.00100 mol/L of _____.
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16. _____ is the concentration of _____ in 0.00100 mol/L of _____.
17. _____ is the concentration of _____ in 0.00100 mol/L of _____.
18. _____ is the concentration of _____ in 0.00100 mol/L of _____.
19. _____ is the concentration of _____ in 0.00100 mol/L of _____.
20. _____ is the concentration of _____ in 0.00100 mol/L of _____.

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19 2 Hydrogen Ions And Acidity Answer Key

John Arthur Wilson



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