

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HBr~~ **hydrobromic acid**
- ~~HI~~ **hydroiodic acid**
- ~~HNO₃~~ **nitric acid**
- ~~HClO₄~~ **perchloric acid**
- ~~Sulfuric hydrosulfate~~ **H₂SO₄**
- ~~potassium hydroxide~~ **KOH**
- ~~lithium hydroxide~~ **LiOH**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~strontium hydroxide~~ **Sr(OH)₂**

HClO₂ is a WEAK acid & does not completely dissociate in water.

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.



3. What is the formula for hydronium? **H₃O⁺**

4. Cu_2CO_3 will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{Cu}_2\text{CO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{CO}_3 + \text{Cu}_2\text{CO}_3^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because Cu_2CO_3 is a weak acid.**
- Would you define Cu_2CO_3 as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is Cu_2CO_3 acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of Cu_2CO_3 : **Cu_2CO_3^-**
- Identify Cu_2CO_3 as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple hydrogens.**

5. F^- will react with water and gain a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{F}^- + \text{H}_2\text{O} \rightarrow \text{HF} + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F^- is a weak base.**
- Would you define F^- as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H^+ ion.**
- In your written equation, is F^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of F^- : **HF**
- Identify F^- as polyprotic, amphoteric, or neither. **neither**

6. HSO_4^- will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{HSO}_4^- + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_4 + \text{SO}_4^{2-}$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HSO_4^- is a weak acid.**
- Would you define HSO_4^- as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is HSO_4^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of HSO_4^- : **SO_4^{2-}**
- Identify HSO_4^- as polyprotic, amphoteric, or neither. **amphoteric, because it has a negative charge (and will like to have a proton) and it has hydrogens (and can lose an H^+).**

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