

# Chemistry Study Guide

Periodic Table of the Elements

1 H Hydrogen	2 He Helium																	18 Ar Argon	19 K Potassium	20 Ca Calcium											36 Kr Krypton	37 Rb Rubidium	38 Sr Strontium											54 Xe Xenon	55 Cs Cesium	56 Ba Barium											86 Rn Radon	87 Fr Francium	88 Ra Radium											118 Og Oganesson
3 Li Lithium	4 Be Beryllium																	12 Mg Magnesium	13 Al Aluminum	14 Si Silicon	15 P Phosphorus	16 S Sulfur	17 Cl Chlorine	18 Ar Argon	39 Y Yttrium	40 Zr Zirconium	41 Nb Niobium	42 Mo Molybdenum	43 Tc Technetium	44 Ru Ruthenium	45 Rh Rhodium	46 Pd Palladium	47 Ag Silver	48 Cd Cadmium	49 In Indium	50 Sn Tin	51 Sb Antimony	52 Te Tellurium	53 I Iodine	54 Xe Xenon	57 La Lanthanum	58 Ce Cerium	59 Pr Praseodymium	60 Nd Neodymium	61 Pm Promethium	62 Sm Samarium	63 Eu Europium	64 Gd Gadolinium	65 Tb Terbium	66 Dy Dysprosium	67 Ho Holmium	68 Er Erbium	69 Tm Thulium	70 Yb Ytterbium	71 Lu Lutetium	89 Ac Actinium	90 Th Thorium	91 Pa Protactinium	92 U Uranium	93 Np Neptunium	94 Pu Plutonium	95 Am Americium	96 Cm Curium	97 Bk Berkelium	98 Cf Californium	99 Es Einsteinium	100 Fm Fermium	101 Md Mendelevium	102 No Nobelium	103 Lr Lawrencium

## Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

## Matter:

Anything that has **mass** and **occupies space** ✎🌐

Examples: air, gases, liquids, solids!

## Atoms:

Atoms are the **smallest, indivisible particle** ✎👉

- **Outer part** → *shell/orbit/energy level* → carries the electrons ( $e^-$ )
- **Inner part** → *nucleus* → houses the protons ( $p^+$ ) and neutrons ( $n$ )
- Outer & Inner part → together these are the **subatomic particles**

## Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	$p^+$	1	+1	Nucleus
Neutron	$n$	1	0	Nucleus
Electron	$e^-$	$\sim 0$	-1	Orbit/Shell

# Section 2 nuclear Chemistry Study Guide Key

**Lei Shi**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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