

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and pK_a and pK_b for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and pK_a and pK_b for weak acids and bases.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-----------------|--------------------------------------|-----------------|
| acid | acid dissociation constant (K_a) | acid ionization |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---------------------------------|----------------------------------|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ | $K_b = \frac{[OH^-][HB^+]}{[B]}$ |
| $pH = -\log[H^+]$ | $pOH = -\log[OH^-]$ |

Part A Completion

Write the appropriate words in the blanks provided. (Use the words in the list on the right.)

1. The K_a of a weak acid is 1.0×10^{-4} . The pK_a of the acid is _____.
2. The K_b of a weak base is 1.0×10^{-4} . The pK_b of the base is _____.
3. The K_a of a weak acid is 1.0×10^{-4} . The pH of a 0.1 M solution of the acid is _____.
4. The K_b of a weak base is 1.0×10^{-4} . The pOH of a 0.1 M solution of the base is _____.
5. The K_a of a weak acid is 1.0×10^{-4} . The pH of a 0.1 M solution of the acid is _____.
6. The K_b of a weak base is 1.0×10^{-4} . The pOH of a 0.1 M solution of the base is _____.
7. The K_a of a weak acid is 1.0×10^{-4} . The pH of a 0.1 M solution of the acid is _____.
8. The K_b of a weak base is 1.0×10^{-4} . The pOH of a 0.1 M solution of the base is _____.

9. K_a
10. K_b
11. pH
12. pOH
13. K_a
14. K_b
15. pH
16. pOH
17. K_a
18. K_b
19. pH
20. pOH

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