



# HYDROGEN IONS AND ACIDITY

## Section Summary

### Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

### Key Vocabulary

<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>	<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>	<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>
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### Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

$$\text{pH} = -\log_{10}[\text{H}^+]$$

### Practice Calculations

Use the following equations to calculate the pH of a solution with a given concentration of hydrogen ions.

**Hydrogen ion concentration:**  $\text{H}^+$  concentration (hydrogen ions)  $\times$  concentration of water molecules  $\times$  concentration of ions in water

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### Answers

1. $\text{pH} = -\log_{10}[\text{H}^+]$	2. $\text{pH} = \log_{10}[\text{OH}^-]$
3. $\text{pH} = -\log_{10}[\text{H}^+]$	4. $\text{pH} = \log_{10}[\text{OH}^-]$
5. $\text{pH} = -\log_{10}[\text{H}^+]$	6. $\text{pH} = \log_{10}[\text{OH}^-]$
7. $\text{pH} = -\log_{10}[\text{H}^+]$	8. $\text{pH} = \log_{10}[\text{OH}^-]$
9. $\text{pH} = -\log_{10}[\text{H}^+]$	10. $\text{pH} = \log_{10}[\text{OH}^-]$
11. $\text{pH} = -\log_{10}[\text{H}^+]$	12. $\text{pH} = \log_{10}[\text{OH}^-]$
13. $\text{pH} = -\log_{10}[\text{H}^+]$	14. $\text{pH} = \log_{10}[\text{OH}^-]$
15. $\text{pH} = -\log_{10}[\text{H}^+]$	16. $\text{pH} = \log_{10}[\text{OH}^-]$
17. $\text{pH} = -\log_{10}[\text{H}^+]$	18. $\text{pH} = \log_{10}[\text{OH}^-]$
19. $\text{pH} = -\log_{10}[\text{H}^+]$	20. $\text{pH} = \log_{10}[\text{OH}^-]$
21. $\text{pH} = -\log_{10}[\text{H}^+]$	22. $\text{pH} = \log_{10}[\text{OH}^-]$
23. $\text{pH} = -\log_{10}[\text{H}^+]$	24. $\text{pH} = \log_{10}[\text{OH}^-]$
25. $\text{pH} = -\log_{10}[\text{H}^+]$	26. $\text{pH} = \log_{10}[\text{OH}^-]$
27. $\text{pH} = -\log_{10}[\text{H}^+]$	28. $\text{pH} = \log_{10}[\text{OH}^-]$
29. $\text{pH} = -\log_{10}[\text{H}^+]$	30. $\text{pH} = \log_{10}[\text{OH}^-]$
31. $\text{pH} = -\log_{10}[\text{H}^+]$	32. $\text{pH} = \log_{10}[\text{OH}^-]$
33. $\text{pH} = -\log_{10}[\text{H}^+]$	34. $\text{pH} = \log_{10}[\text{OH}^-]$
35. $\text{pH} = -\log_{10}[\text{H}^+]$	36. $\text{pH} = \log_{10}[\text{OH}^-]$
37. $\text{pH} = -\log_{10}[\text{H}^+]$	38. $\text{pH} = \log_{10}[\text{OH}^-]$
39. $\text{pH} = -\log_{10}[\text{H}^+]$	40. $\text{pH} = \log_{10}[\text{OH}^-]$
41. $\text{pH} = -\log_{10}[\text{H}^+]$	42. $\text{pH} = \log_{10}[\text{OH}^-]$
43. $\text{pH} = -\log_{10}[\text{H}^+]$	44. $\text{pH} = \log_{10}[\text{OH}^-]$
45. $\text{pH} = -\log_{10}[\text{H}^+]$	46. $\text{pH} = \log_{10}[\text{OH}^-]$
47. $\text{pH} = -\log_{10}[\text{H}^+]$	48. $\text{pH} = \log_{10}[\text{OH}^-]$
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77. $\text{pH} = -\log_{10}[\text{H}^+]$	78. $\text{pH} = \log_{10}[\text{OH}^-]$
79. $\text{pH} = -\log_{10}[\text{H}^+]$	80. $\text{pH} = \log_{10}[\text{OH}^-]$
81. $\text{pH} = -\log_{10}[\text{H}^+]$	82. $\text{pH} = \log_{10}[\text{OH}^-]$
83. $\text{pH} = -\log_{10}[\text{H}^+]$	84. $\text{pH} = \log_{10}[\text{OH}^-]$
85. $\text{pH} = -\log_{10}[\text{H}^+]$	86. $\text{pH} = \log_{10}[\text{OH}^-]$
87. $\text{pH} = -\log_{10}[\text{H}^+]$	88. $\text{pH} = \log_{10}[\text{OH}^-]$
89. $\text{pH} = -\log_{10}[\text{H}^+]$	90. $\text{pH} = \log_{10}[\text{OH}^-]$
91. $\text{pH} = -\log_{10}[\text{H}^+]$	92. $\text{pH} = \log_{10}[\text{OH}^-]$
93. $\text{pH} = -\log_{10}[\text{H}^+]$	94. $\text{pH} = \log_{10}[\text{OH}^-]$
95. $\text{pH} = -\log_{10}[\text{H}^+]$	96. $\text{pH} = \log_{10}[\text{OH}^-]$
97. $\text{pH} = -\log_{10}[\text{H}^+]$	98. $\text{pH} = \log_{10}[\text{OH}^-]$
99. $\text{pH} = -\log_{10}[\text{H}^+]$	100. $\text{pH} = \log_{10}[\text{OH}^-]$

# 19 2 Hydrogen Ions And Acidity Work Sheet Answers

**Brendan G. Carr**

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