



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Define acid and base and relate them to the pH of a solution.
- Define strong and weak acids.

Key Vocabulary

| | |
|---|---|
| <ul style="list-style-type: none">hydrogen ionacidbasestrong acidweak acidpH | <ul style="list-style-type: none">hydrogen ion concentrationacidicbasicneutraltitrationtitration curve |
|---|---|

Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

Key Diagrams

- Titration curve
- Titration curve

Points of Consideration

Acidic and basic properties are determined by the concentration of hydrogen ions in a solution. The concentration of hydrogen ions is measured by the pH of a solution.

Hydrogen ions are the only ions that contribute to the concentration of ions in a solution. The concentration of ions in a solution is determined by the concentration of hydrogen ions.

Hydrogen ions are the only ions that contribute to the concentration of ions in a solution. The concentration of ions in a solution is determined by the concentration of hydrogen ions.

Hydrogen ions are the only ions that contribute to the concentration of ions in a solution. The concentration of ions in a solution is determined by the concentration of hydrogen ions.

Key Equations

| | |
|--------------------------------------|--------------------------------------|
| $\text{pH} = -\log_{10}[\text{H}^+]$ | $\text{pH} = \log_{10}[\text{OH}^-]$ |
| $\text{pH} = 7$ | $\text{pH} = 14$ |
| $\text{pH} = 1$ | $\text{pH} = 13$ |
| $\text{pH} = 2$ | $\text{pH} = 12$ |
| $\text{pH} = 3$ | $\text{pH} = 11$ |
| $\text{pH} = 4$ | $\text{pH} = 10$ |
| $\text{pH} = 5$ | $\text{pH} = 9$ |
| $\text{pH} = 6$ | $\text{pH} = 8$ |
| $\text{pH} = 7$ | $\text{pH} = 7$ |
| $\text{pH} = 8$ | $\text{pH} = 6$ |
| $\text{pH} = 9$ | $\text{pH} = 5$ |
| $\text{pH} = 10$ | $\text{pH} = 4$ |
| $\text{pH} = 11$ | $\text{pH} = 3$ |
| $\text{pH} = 12$ | $\text{pH} = 2$ |
| $\text{pH} = 13$ | $\text{pH} = 1$ |
| $\text{pH} = 14$ | $\text{pH} = 0$ |

19 2 Hydrogen Ions And Acidity Worksheet Answers

Sir Tarun Rupani

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