

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, given  $K_a$  or  $K_b$ .
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, given  $K_a$  or  $K_b$ .
3. Calculate the pH of weak acid and base solutions.

## Vocabulary

- |                            |                   |              |
|----------------------------|-------------------|--------------|
| acid dissociation constant | 1. conjugate base | 2. weak acid |
| acid ionization            | 3. conjugate acid | 3. weak base |
| acid ionization constant   | 4. conjugate acid | 4. weak base |

## Key Equations

- |                                    |                                   |
|------------------------------------|-----------------------------------|
| 1. $K_a = \frac{[H^+][A^-]}{[HA]}$ | 2. $K_b = \frac{[HB][OH^-]}{[B]}$ |
| 3. $K_a \cdot K_b = K_w$           | 4. $K_a \cdot K_b = K_w$          |

## Part A Completion

Write the appropriate words in the blanks to complete the following sentences. Write the chemical formula of the acid or base in the blank.

1. The  $K_a$  of acetic acid is  $1.8 \times 10^{-5}$ .
2. The  $K_b$  of ammonia is  $1.8 \times 10^{-5}$ .
3. The  $K_a$  of acetic acid is  $1.8 \times 10^{-5}$ .
4. The  $K_b$  of ammonia is  $1.8 \times 10^{-5}$ .
5. The  $K_a$  of acetic acid is  $1.8 \times 10^{-5}$ .
6. The  $K_b$  of ammonia is  $1.8 \times 10^{-5}$ .
7. The  $K_a$  of acetic acid is  $1.8 \times 10^{-5}$ .
8. The  $K_b$  of ammonia is  $1.8 \times 10^{-5}$ .
9. The  $K_a$  of acetic acid is  $1.8 \times 10^{-5}$ .
10. The  $K_b$  of ammonia is  $1.8 \times 10^{-5}$ .

1. acetic acid
2.  $1.8 \times 10^{-5}$
3.  $1.8 \times 10^{-5}$
4.  $1.8 \times 10^{-5}$
5.  $1.8 \times 10^{-5}$
6.  $1.8 \times 10^{-5}$
7.  $1.8 \times 10^{-5}$
8.  $1.8 \times 10^{-5}$
9.  $1.8 \times 10^{-5}$
10.  $1.8 \times 10^{-5}$

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**Raffaella Di Napoli**



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