

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and pK_a and pK_b for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and pK_a and pK_b for weak acids and bases.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-----------------|--------------------------------------|-----------------|
| acid | acid dissociation constant (K_a) | acid ionization |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

Part A Completion

Write the chemical formula for each of the following acids, bases, and salts. Identify the acid, base, or salt.

1. Hydrochloric acid, HCl , (Equation 19.1)

2. Acetic acid, CH_3COOH , (Equation 19.1)

3. Nitric acid, HNO_3 , (Equation 19.1)

4. Sulfuric acid, H_2SO_4 , (Equation 19.1)

5. Hydrofluoric acid, HF , (Equation 19.1)

6. Hydrobromic acid, HBr , (Equation 19.1)

7. Hydroiodic acid, HI , (Equation 19.1)

8. Hydrocyanic acid, HCN , (Equation 19.1)

9. Hydroperfluoroboric acid, HF_2 , (Equation 19.1)

10. Hydroperfluoroboric acid, HF_2 , (Equation 19.1)

11. Hydroperfluoroboric acid, HF_2 , (Equation 19.1)

1. Hydrochloric acid
2. Acetic acid
3. Nitric acid
4. Sulfuric acid
5. Hydrofluoric acid
6. Hydrobromic acid
7. Hydroiodic acid
8. Hydrocyanic acid
9. Hydroperfluoroboric acid
10. Hydroperfluoroboric acid
11. Hydroperfluoroboric acid

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Valerie C. Scanlon, Tina Sanders



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