

ACS Final Exam Practice 1

The exam is 70 questions in 110 minutes. You may use the equation sheet and periodic table but no other materials

Section 1 Foundations

PQ-1. The speed of light is  $3.0 \times 10^8 \text{ m}\cdot\text{s}^{-1}$ . What is this speed in  $\text{nm}\cdot\text{ms}^{-1}$ ?

- (A)  $3.0 \times 10^{-1} \text{ nm}\cdot\text{ms}^{-1}$       (B)  $3.0 \times 10^2 \text{ nm}\cdot\text{ms}^{-1}$       (C)  $3.0 \times 10^{14} \text{ nm}\cdot\text{ms}^{-1}$       (D)  $3.0 \times 10^{23} \text{ nm}\cdot\text{ms}^{-1}$

PQ-4. What is the correctly reported mass of water based on the data in the table?

Mass of beaker and water	29.62 g
Mass of beaker only	28.3220 g

- (A) 1.3 g      (B) 1.30 g      (C) 1.298 g      (D) 1.2980 g

PQ-5. What is the name of  $\text{Ti}_2(\text{PO}_4)_3$ ?

- (A) titanium phosphate      (B) titanium(III) phosphate  
(C) titanium(IV) phosphate      (D) titanium tetraphosphate

PQ-6. The formula of strontium hexafluorosilicate is  $\text{SrSiF}_6$ . What is the formula of aluminum hexafluorosilicate?

- (A)  $\text{AlSiF}_6$       (B)  $\text{Al}_2(\text{SiF}_6)_3$       (C)  $\text{Al}_2\text{SiF}_6$       (D)  $\text{Al}_2(\text{SiF}_6)_2$

PQ-8. What is the correct name for  $\text{N}_2\text{O}_5$ ?

- (A) nitrogen(II) oxide      (B) nitrogen(V) oxide  
(C) dinitrogen oxide      (D) dinitrogen pentoxide

PQ-12. If a palladium nanoparticle has a density of  $12.0 \text{ g}\cdot\text{cm}^{-3}$ , what is the mass of a nanoparticle with a volume of  $1.84 \times 10^{-23} \text{ cm}^3$ ?

- (A)  $2.21 \times 10^{-23} \text{ g}$       (B)  $1.63 \times 10^{-20} \text{ g}$       (C)  $6.52 \times 10^{-21} \text{ g}$       (D)  $1.53 \times 10^{-22} \text{ g}$

# 70 Question Chemistry Final Exam Answers

**J Spring**



## 70 Question Chemistry Final Exam Answers:

**Encyclopedia of Chemical Processing and Design** John J. McKetta Jr,1993-06-02 Written by engineers for engineers with over 150 International Editorial Advisory Board members this highly lauded resource provides up to the minute information on the chemical processes methods practices products and standards in the chemical and related industries

*Lessons in Organic Chemistry* G. S. Turpin,1894     **Inorganic Chemistry for Beginners** Henry Enfield Roscoe,Joseph Lunt,1897     *The Publishers Weekly* ,1892     **Science Reporter** ,1994     Boston Journal of Chemistry and Popular Science Review ,1881     *Report Upon the Survey of the University of Wisconsin* Wisconsin. State Board of Public Affairs,1914     *A Class-book of Inorganic Chemistry; with tables of chemical analysis, etc* David MORRIS (B.A.),1870

**Index to the British Catalogue of Books** ,1876     **The American Bookseller** ,1886     **The Chemical News and Journal of Physical Science** ,1881     *Elementary Physical Geography* Ralph Stockman Tarr,1895     **Joint Computer Conference** ,1971     Grade Booster for CfE SQA Exam Revision - Higher Chemistry Grade Booster for SQA Exam Revision: Maximise Marks and Minimise Mistakes to Achieve Your Best Possible Mark Tom Speirs,Leckie,2022-02-11 Exam Board SQA Level Higher Subject Chemistry First Teaching 2014 First Exam 2015 This CfE Higher Chemistry Grade Booster is the essential guide to exam skills It includes detailed advice on how to approach and answer the different types of question you will find in the exam and has been written by an experienced teacher and exam expert     Metals and Materials ,1973

*Indian Book Industry* ,1999     *How to Pass National Teacher Examinations* College Publishing Corporation, Brooklyn,1967     GRE ... Registration & Information Bulletin ,1992     *Principles of Chemistry* Joseph Hiram Roe,1929  
*Papers and Discussions Presented* ,1971

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