

Acids & Bases HOMEWORK

Homework

1. Propose the correct ion search of the strong acid listed below. Propose the incorrect ion search of the strong base listed below.

- a. **HF**: **Hydrogen fluoride**
- b. **HI**: **Hydrogen iodide**
- c. **HNO₃**: **nitric acid**
- d. **HClO₄**: **perchloric acid**
- e. **HBrO₃**: **hypobromous acid**
- f. **HClO₄**: **chloric acid**
- g. **H₂SO₄**: **hydrochloric acid**
- h. **H₂SO₄**: **hydrobromic acid**
- i. **H₂SO₄**: **hydroiodic acid**
- j. **H₂SO₄**: **hydrochloric acid**
- k. **H₂SO₄**: **hydrobromic acid**
- l. **H₂SO₄**: **hydroiodic acid**

2. Off the acids and bases listed below, write the ones that completely dissociate in water.

| Acid | Base | Mixture | Neutral | Acidic | Basic | Neutral | Basic |
|------|------|---------|---------|--------|-------|---------|-------|
|------|------|---------|---------|--------|-------|---------|-------|

3. Write the dissociation for **H₂SO₄**.

4. **HClO₄** will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{H}_3\text{ClO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_2\text{ClO}_4 + \text{H}_3\text{O}^+$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating H_3ClO_4 in a acidic acid.

c. Dissolve **1.00\text{M}** NaClO_4 in a portion of water and a portion of **1.00\text{M}** **NaOH**. Dissolve **1.00\text{M}** H_3O^+ in a portion of water.

d. In your **reaction** prediction, is HClO_4 acting as a **Brønsted-Lowry** acid or a **Brønsted-Lowry** base? **Brønsted-Lowry acid**. Because HClO_4 is a **strong** acid.

e. Identify the conjugate base of HClO_4 : **ClO_4^-**

f. Identify NaClO_4 as **strong**, **weak**, **neutral**, **acidic**, **basic**, **neutral**, **acidic**, **basic**, **neutral**, **acidic**, **basic**.

5. H^+ will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{H}_3\text{O}^+ + \text{H}_2\text{O} \longrightarrow \text{H}_2\text{O}^+ + \text{H}_2\text{O}$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating H_3O^+ in a acidic base.

c. Dissolve **1.00\text{M}** H_3O^+ in a portion of water and a portion of **1.00\text{M}** **NaOH**. Dissolve **1.00\text{M}** H_2O^+ in a portion of water.

d. In your **reaction** prediction, is H_3O^+ acting as a **Brønsted-Lowry** acid or a **Brønsted-Lowry** base? **Brønsted-Lowry acid**. Because H_3O^+ is a **strong** acid.

e. Identify the conjugate base of H_3O^+ : **H_2O**

f. Identify H_3O^+ as **strong**, **weak**, **neutral**, **acidic**, **basic**, **neutral**, **acidic**, **basic**.

6. **H₂SO₄** will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_2\text{SO}_4^+ + \text{H}_3\text{O}^+$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating H_2SO_4 in a acidic acid.

c. Dissolve **1.00\text{M}** NaHSO_4 in a portion of water and a portion of **1.00\text{M}** **NaOH**. Dissolve **1.00\text{M}** H_2O^+ in a portion of water.

d. In your **reaction** prediction, is H_2SO_4 acting as a **Brønsted-Lowry** acid or a **Brønsted-Lowry** base? **Brønsted-Lowry acid**. Because H_2SO_4 is a **strong** acid.

e. Identify the conjugate base of H_2SO_4 : **HSO_4^-**

f. Identify NaHSO_4 as **strong**, **weak**, **neutral**, **acidic**, **basic**, **neutral**, **acidic**, **basic**. **strong** acid. Because NaHSO_4 is a **strong** acid.

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Charles Kendall Adams

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