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buffers typically consist of an acid base pair with the acid and base differing by the presence or absence of a proton a conjugate acid base pair for instance one of the buffers that maintain the ph of human blood involves carbonic acid H_2CO_3 and its conjugate base the bicarbonate ion HCO_3^-

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an acid and a base react not to produce a salt and a solvent but to form a new acid and a new base the concept of neutralization is thus absent brønsted lowry acid base behavior is formally independent of any solvent making it more all encompassing than the arrhenius model

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[a binary acid](#) is an acid that consists of hydrogen and one other element the most common binary acids contain a halogen the acid name begins with the prefix hydro followed by the base name of the anion followed by the suffix ic

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magnesium aluminium or iron acids have a low pH value i.e. below 7 salts and water are produced when an acid reacts with base or alkalis most of the acids are solids

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