

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode $T \gg \theta_M$) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute $\frac{3}{2}R$ to the molar heat capacity. If we denote the number of active rotational modes by ν_R^\ddagger (so for most molecules at normal temperatures $\nu_R^\ddagger = 2$ for linear molecules, and 3 for non-linear molecules), then

the total heat capacity is given by $C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \frac{\nu_V}{2}R$, where ν_V is the number of active vibrational modes. In the equipartition limit $\nu_V = 3N - \nu_R^\ddagger - 3$, so that $C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \frac{3N - \nu_R^\ddagger - 3}{2}R = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \frac{3N}{2}R - \frac{\nu_R^\ddagger}{2}R - \frac{3}{2}R = \frac{3N}{2}R$.

$$C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \frac{3N}{2}R$$

The equipartition value of $C_{V,m}$ is $\frac{3}{2}R$ for monatomic gases, $\frac{5}{2}R$ for linear diatomic gases, and $\frac{6}{2}R$ for non-linear triatomic gases. For polyatomic gases, the equipartition value of $C_{V,m}$ is $\frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \frac{3N}{2}R$, where ν_R^\ddagger is the number of active rotational modes. For linear molecules, $\nu_R^\ddagger = 2$, and for non-linear molecules, $\nu_R^\ddagger = 3$. The equipartition value of $C_{V,m}$ is $\frac{5}{2}R$ for linear diatomic gases, and $\frac{6}{2}R$ for non-linear triatomic gases. For polyatomic gases, the equipartition value of $C_{V,m}$ is $\frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \frac{3N}{2}R$, where ν_R^\ddagger is the number of active rotational modes. For linear molecules, $\nu_R^\ddagger = 2$, and for non-linear molecules, $\nu_R^\ddagger = 3$.

Solutions to exercises

The thermal energy is given by $U = \frac{1}{2}k_B T$ per degree of freedom. When equipartition is valid, the thermal energy is $U = \frac{1}{2}k_B T$ per degree of freedom.

$$U = \frac{1}{2}k_B T = \frac{1}{2}(1.38 \times 10^{-23} \text{ J K}^{-1})(300 \text{ K}) = 2.07 \times 10^{-21} \text{ J}$$

For a diatomic molecule, the thermal energy is $U = \frac{5}{2}k_B T$.

$$U = \frac{5}{2}k_B T = \frac{5}{2}(1.38 \times 10^{-23} \text{ J K}^{-1})(300 \text{ K}) = 1.03 \times 10^{-20} \text{ J}$$

For a triatomic molecule,

$$U = \frac{6}{2}k_B T = \frac{6}{2}(1.38 \times 10^{-23} \text{ J K}^{-1})(300 \text{ K}) = 1.22 \times 10^{-20} \text{ J}$$

$$C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \frac{3N}{2}R$$

For a diatomic molecule, $\nu_R^\ddagger = 2$.

Physical Chemistry 9th Edition Instructors Solution Guide

Michael Hoelscher



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