

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the  $K_a$  and  $K_b$  of acids and bases and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                   |                             |                  |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain                | 6. acid rain     |
| 7. autoionization | 8. acid rain                | 9. acid          |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the equation and/or the numerical value(s) for the reaction(s) described. Show the calculation for pH or pOH. Write the chemical formula for the acid or base.

1. Write the equation for the autoionization of water.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
2. Write the equation for the autoionization of water.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
3. Write the equation for the autoionization of water.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
4. Write the equation for the autoionization of water.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
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1.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 2.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 3.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 4.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 5.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 6.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 7.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 8.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 9.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 10.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 11.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 12.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 13.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 14.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 15.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 16.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 17.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 18.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 19.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 20.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$

# 19 2 Hydrogen Ions And Acidity Worksheet Answers

**Advisory Council on College Chemistry**



## 19 2 Hydrogen Ions And Acidity Worksheet Answers:

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**Hydrogen Ions** Hubert Thomas Stanley Britton,1943      The ABC of Hydrogen Ion Control La Motte Chemical Products Company,William Anthony Taylor,1928

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