

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-----------------|--------------------------------------|----------------|
| acid | acid dissociation constant (K_a) | base |
| acidic solution | base dissociation constant (K_b) | basic solution |
| | | |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

Part A Completion

Write the appropriate words in the blanks to complete the following sentences. Write the words in the blanks.

1. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
2. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
3. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
4. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
5. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
6. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
7. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
8. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .

9. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
10. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
11. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
12. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
13. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
14. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
15. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
16. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .

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web in today's experiment you will determine the solubility product K_{sp} of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the K_{sp} of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq})$ the solubility product expression describes in mathematical terms the equilibrium

[Ksp for calcium hydroxide university of manitoba](#) - Dec 31 2022

web K_{sp} for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a pH regulating agent notice that calcium hydroxide is divalent and thus has twice

Ksp calcium hydroxide chem132 determination of the K_{sp} of calcium - Jun 24 2022

web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $\text{Ca(OH)}_2 \rightarrow \text{Ca}^{2+} + 2\text{OH}^{-}$ $5 \times 10^{-2} \text{ M}$ calcium ions $5 \times 10^{-2} \text{ M}$ $2 \times 10^{-2} \text{ M}$ hydroxide ions calculate calcium hydroxide K_{sp} using the concentrations of calcium ions and hydroxide ions

solubility of calcium hydroxide wpmu dev - Oct 29 2022

web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the K_{sp} of a slightly soluble basic salt and how solubility changes with temperature K_{sp} is another form of equilibrium expression that deals with solubility the reaction that will take place is $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^{-}$

solved experiment 8 determining K_{sp} for calcium hydroxide - Aug 27 2022

web question experiment 8 determining K_{sp} for calcium hydroxide post lab report sheet data table 2 points trial 1 trial 2 HCl molarity 0.050 M volume of filtered Ca(OH)_2 solution mL 5.00 5.00 3.71 13.75 equivalence point volume mL OH⁻ in saturated Ca(OH)_2 Ca^{2+} in saturated Ca(OH)_2 K_{sp} of Ca(OH)_2 average K_{sp} of Ca(OH)_2

[chemistry 1210 lab 11 experiment 11 determining the \$K_{sp}\$ of calcium](#) - Sep 08 2023

web the chemical equation used was $\text{Ca(OH)}_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq})$ the solubility product was determined to be $K_{sp} = 5.02 \times 10^{-6}$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

18 1 solubility product constant ksp chemistry libretexts - Feb 01 2023

web jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3(\text{PO}_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K_{\text{sp}} = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2$

post lab report 8 determining the ksp of calcium hydroxide - Jul 06 2023

web lab report 8 determining the ksp of calcium hydroxide the purpose of this was to titrate a saturated $\text{Ca}(\text{OH})_2$ solution with a standard HCl solution and find the ksp or the OH^- concentration was found and then the concentration was found and used for this experiment a titration of calcium hydroxide with hydrochloric acid

solved determining the ksp of calcium hydroxide - Apr 22 2022

web question determining the ksp of calcium hydroxide $\text{Ca}(\text{OH})_2 \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{OH}^-(\text{aq})$ for this experiment the titration of calcium hydroxide with hydrochloric acid was investigated determining the ksp of calcium hydroxide $\text{Ca}(\text{OH})_2 \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{OH}^-(\text{aq})$ for this experiment the titration of calcium

solved chem 182 experiment 8 determining the ksp of calcium chegg - Sep 27 2022

web question chem 182 experiment 8 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water solution of $\text{Ca}(\text{OH})_2$ is represented in equation form solution of calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution as shown below

determination of the ksp of calcium hydroxide autosaved - Jun 05 2023

web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for $\text{Ca}(\text{OH})_2$ was the determination through pH method 2 introduction calcium hydroxide is a strong base soluble in water

exp 4 determining the solubility product constant ksp of calcium - Jul 26 2022

web plasma membrane homework symbiosis practice worksheet benzoin condensation grade a determining the ksp of calcium hydroxide after reading the lab complete items a b c and d title purpose chemicals and equipment and summary of procedure as described on page 10 of exp 1 on an 8 1/2 x 11 sheet of paper

lab 9 determination of the solubility product ksp chegg com - May 24 2022

web question lab 9 determination of the solubility product ksp of calcium hydroxide introduction ionic compounds that are classified as insoluble based on solubility rules are actually slightly soluble each of these insoluble compounds actually dissolves to

experimental determination of the solubility product of calcium hydroxide - Nov 29 2022

web we will filter the solid calcium hydroxide from the solution and then titrate the dissolved calcium hydroxide in solution

with a known concentration of hydrochloric acid using titrimetric calculations we will determine the Ca^{2+} and OH^{-} concentrations in solution

ksp of calcium hydroxide lab report for determining the ksp of - Oct 09 2023

web calculate the ksp for calcium hydroxide $\text{ksp Ca}^{2+} \text{OH}^{-}$ 2.0×10^{-2} ksp 5×10^{-4} find the accepted value of the ksp for calcium hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error
accepted value for ksp for sodium hydroxide is 5×10^{-6}