

## 19.5

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the  $K_a$  and  $K_b$  values and vice versa.
3. Calculate the pH of weak acids.

## Vocabulary

- |                   |                             |                   |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization | 5. $K_a$ value              | 6. $K_b$ value    |
| 7. ionization     | 8. $K_a$ value              | 9. $K_b$ value    |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the equation and/or the numerical value(s) of the values indicated. Use the information in the section above to complete each question.

1. What is the  $\text{pH}$  of a  $0.1 \text{ M}$  solution of hydrochloric acid?
2. What is the  $\text{pH}$  of a  $0.1 \text{ M}$  solution of sodium hydroxide?
3. What is the  $\text{pH}$  of a  $0.1 \text{ M}$  solution of acetic acid?
4. What is the  $\text{pH}$  of a  $0.1 \text{ M}$  solution of ammonia?
5. What is the  $\text{pH}$  of a  $0.1 \text{ M}$  solution of hydrofluoric acid?
6. What is the  $\text{pH}$  of a  $0.1 \text{ M}$  solution of hydrocyanic acid?
7. What is the  $\text{pH}$  of a  $0.1 \text{ M}$  solution of hydrobromic acid?
8. What is the  $\text{pH}$  of a  $0.1 \text{ M}$  solution of hydroiodic acid?
9. What is the  $\text{pH}$  of a  $0.1 \text{ M}$  solution of hydrochloric acid?
10. What is the  $\text{pH}$  of a  $0.1 \text{ M}$  solution of hydrobromic acid?
11. What is the  $\text{pH}$  of a  $0.1 \text{ M}$  solution of hydroiodic acid?

12. What is the  $\text{pH}$  of a  $0.1 \text{ M}$  solution of hydrochloric acid?
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20. What is the  $\text{pH}$  of a  $0.1 \text{ M}$  solution of hydroiodic acid?

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