



# HYDROGEN IONS AND ACIDITY

## Section Summary

### Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

### Key Vocabulary

<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>	<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>	<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>
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### Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

$$\text{pH} = -\log_{10}[\text{H}^+]$$

### Practice Calculations

Use the following equations to calculate the pH of a solution with a hydrogen ion concentration of  $1.0 \times 10^{-4} \text{ M}$ .

$\text{pH} = -\log_{10}[\text{H}^+]$        $\text{pH} = -\log_{10}[\text{H}^+]$   
concentration      concentration

Concentration	pH
$1.0 \times 10^{-4} \text{ M}$	4.0
$1.0 \times 10^{-5} \text{ M}$	5.0
$1.0 \times 10^{-6} \text{ M}$	6.0
$1.0 \times 10^{-7} \text{ M}$	7.0
$1.0 \times 10^{-8} \text{ M}$	8.0
$1.0 \times 10^{-9} \text{ M}$	9.0
$1.0 \times 10^{-10} \text{ M}$	10.0
$1.0 \times 10^{-11} \text{ M}$	11.0
$1.0 \times 10^{-12} \text{ M}$	12.0
$1.0 \times 10^{-13} \text{ M}$	13.0
$1.0 \times 10^{-14} \text{ M}$	14.0

## 19 2 Hydrogen Ions And Acidity Section Review Answers

**Linda Swisher, Gary A. Thibodeau**

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**Journal of the American Medical Association** American Medical Association,1924 Includes proceedings of the association papers read at the annual sessions and lists of current medical literature     Addison-Wesley Chemistry ,1990

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**Electrochemical and Metallurgical Industry** Eugene Franz Roeber,Howard Coon Parmelee,1920     *ASTM Bulletin* American Society for Testing Materials,1956     Journal of the Chemical Society Chemical Society (Great Britain),1950 Titles of chemical papers in British and foreign journals included in Quarterly journal v 1-12     **Study Guide to Accompany**

**Structure and Function of the Body** Linda Swisher,Gary A. Thibodeau,2003-11 This comprehensive resource provides a variety of exercises for readers to apply and test their knowledge It contains matching fill in the blanks crossword puzzles word find unscramble the word application questions diagrams and page number references in the answer key December 2003

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**Bibliography of Agriculture** ,1982-12     *Experiments and Exercises in Basic Chemistry* Steven Murov,Brian Stedjee,2000-02-03     *The Lancet* ,1947     New York Produce Review and American Creamery ,1936     **Crop Physiology Abstracts** ,1985     **Forensic Science Review** ,2001     **Comprehensive Review**

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