



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

Key Vocabulary

<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base
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Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

$$\text{pH} = -\log_{10}[\text{H}^+]$$

Practice Calculations

Use the following equations to calculate the pH of a solution with a given concentration of hydrogen ions.

Hydrogen ion concentration: H^+ concentration (hydrogen ions) \times concentration of other ions \times concentration of water \times concentration of other ions

Hydrogen ion concentration: H^+ concentration (hydrogen ions) \times concentration of other ions \times concentration of water \times concentration of other ions \times concentration of other ions

Hydrogen ion concentration: H^+ concentration (hydrogen ions) \times concentration of other ions \times concentration of water \times concentration of other ions \times concentration of other ions

Answers

1. $\text{pH} = -\log_{10}[\text{H}^+]$	2. $\text{pH} = \log_{10}[\text{OH}^-]$
3. $\text{pH} = -\log_{10}[\text{H}^+]$	4. $\text{pH} = \log_{10}[\text{OH}^-]$
5. $\text{pH} = -\log_{10}[\text{H}^+]$	6. $\text{pH} = \log_{10}[\text{OH}^-]$
7. $\text{pH} = -\log_{10}[\text{H}^+]$	8. $\text{pH} = \log_{10}[\text{OH}^-]$
9. $\text{pH} = -\log_{10}[\text{H}^+]$	10. $\text{pH} = \log_{10}[\text{OH}^-]$
11. $\text{pH} = -\log_{10}[\text{H}^+]$	12. $\text{pH} = \log_{10}[\text{OH}^-]$
13. $\text{pH} = -\log_{10}[\text{H}^+]$	14. $\text{pH} = \log_{10}[\text{OH}^-]$
15. $\text{pH} = -\log_{10}[\text{H}^+]$	16. $\text{pH} = \log_{10}[\text{OH}^-]$
17. $\text{pH} = -\log_{10}[\text{H}^+]$	18. $\text{pH} = \log_{10}[\text{OH}^-]$
19. $\text{pH} = -\log_{10}[\text{H}^+]$	20. $\text{pH} = \log_{10}[\text{OH}^-]$
21. $\text{pH} = -\log_{10}[\text{H}^+]$	22. $\text{pH} = \log_{10}[\text{OH}^-]$
23. $\text{pH} = -\log_{10}[\text{H}^+]$	24. $\text{pH} = \log_{10}[\text{OH}^-]$
25. $\text{pH} = -\log_{10}[\text{H}^+]$	26. $\text{pH} = \log_{10}[\text{OH}^-]$
27. $\text{pH} = -\log_{10}[\text{H}^+]$	28. $\text{pH} = \log_{10}[\text{OH}^-]$
29. $\text{pH} = -\log_{10}[\text{H}^+]$	30. $\text{pH} = \log_{10}[\text{OH}^-]$
31. $\text{pH} = -\log_{10}[\text{H}^+]$	32. $\text{pH} = \log_{10}[\text{OH}^-]$
33. $\text{pH} = -\log_{10}[\text{H}^+]$	34. $\text{pH} = \log_{10}[\text{OH}^-]$
35. $\text{pH} = -\log_{10}[\text{H}^+]$	36. $\text{pH} = \log_{10}[\text{OH}^-]$
37. $\text{pH} = -\log_{10}[\text{H}^+]$	38. $\text{pH} = \log_{10}[\text{OH}^-]$
39. $\text{pH} = -\log_{10}[\text{H}^+]$	40. $\text{pH} = \log_{10}[\text{OH}^-]$
41. $\text{pH} = -\log_{10}[\text{H}^+]$	42. $\text{pH} = \log_{10}[\text{OH}^-]$
43. $\text{pH} = -\log_{10}[\text{H}^+]$	44. $\text{pH} = \log_{10}[\text{OH}^-]$
45. $\text{pH} = -\log_{10}[\text{H}^+]$	46. $\text{pH} = \log_{10}[\text{OH}^-]$
47. $\text{pH} = -\log_{10}[\text{H}^+]$	48. $\text{pH} = \log_{10}[\text{OH}^-]$
49. $\text{pH} = -\log_{10}[\text{H}^+]$	50. $\text{pH} = \log_{10}[\text{OH}^-]$
51. $\text{pH} = -\log_{10}[\text{H}^+]$	52. $\text{pH} = \log_{10}[\text{OH}^-]$
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87. $\text{pH} = -\log_{10}[\text{H}^+]$	88. $\text{pH} = \log_{10}[\text{OH}^-]$
89. $\text{pH} = -\log_{10}[\text{H}^+]$	90. $\text{pH} = \log_{10}[\text{OH}^-]$
91. $\text{pH} = -\log_{10}[\text{H}^+]$	92. $\text{pH} = \log_{10}[\text{OH}^-]$
93. $\text{pH} = -\log_{10}[\text{H}^+]$	94. $\text{pH} = \log_{10}[\text{OH}^-]$
95. $\text{pH} = -\log_{10}[\text{H}^+]$	96. $\text{pH} = \log_{10}[\text{OH}^-]$
97. $\text{pH} = -\log_{10}[\text{H}^+]$	98. $\text{pH} = \log_{10}[\text{OH}^-]$
99. $\text{pH} = -\log_{10}[\text{H}^+]$	100. $\text{pH} = \log_{10}[\text{OH}^-]$

19 2 Hydrogen Ions And Acidity Work Sheet Answers

J Spring

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