

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid rain |
| 7. autoionization | 8. acid rain | 9. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the unknown quantity that are indicated in the problem. Be sure to include units whenever a unit is appropriate.

1. What substance is _____, which hydrogen ion concentration is _____? The concentration of this acid is _____.
2. What substance is _____, which hydrogen ion concentration is _____? The concentration of this acid is _____.
3. What substance is _____, which hydrogen ion concentration is _____? The concentration of this acid is _____.
4. What substance is _____, which hydrogen ion concentration is _____? The concentration of this acid is _____.
5. What substance is _____, which hydrogen ion concentration is _____? The concentration of this acid is _____.
6. What substance is _____, which hydrogen ion concentration is _____? The concentration of this acid is _____.
7. What substance is _____, which hydrogen ion concentration is _____? The concentration of this acid is _____.

8. _____
9. _____
10. _____
11. _____
12. _____
13. _____
14. _____
15. _____
16. _____
17. _____
18. _____

19 2 Hydrogen Ions And Acidity Work Sheet Answers

J Spring



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