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Chapter 7

Equilibria for Multiple Component Systems

7.2. With only a single component and phase, if only $\text{Fe}_2(\text{SO}_4)_3$ were present in our system, we would have $C - P + 2 = 1 - 1 + 2 = 2$ degrees of freedom.

7.4. For a single-component system, there will never be negative degrees of freedom. The maximum number of phases that can exist in equilibrium is three, so the Gibbs phase rule would yield $1 - 3 + 2 = 0$ degrees of freedom (like might occur at the triple point). For a two-component system, the maximum number of phases that can exist in equilibrium is four. There will never be negative degrees of freedom.

7.6. In this reaction, there are only three independent components (the amount of the fourth can be determined from the equilibrium constant) and 3 phases, two solids and a gas mixture. Therefore, there are $3 - 3 + 2 = 2$ degrees of freedom.

7.8. For water: $n = \frac{PV}{RT} = \frac{(23.76 \text{ torr})(5.00 \text{ L})}{(0.08205 \text{ L} \cdot \text{atm/mol} \cdot \text{K})(298.15 \text{ K})} \times \frac{1 \text{ atm}}{760 \text{ torr}} = 6.39 \times 10^{-3}$ mol of water is needed to ensure that there is a liquid and gas phase. This equals 0.115 grams.
For methanol:
$$n = \frac{PV}{RT} = \frac{(125.0 \text{ torr})(5.00 \text{ L})}{(0.08205 \text{ L} \cdot \text{atm/mol} \cdot \text{K})(298.15 \text{ K})} \times \frac{1 \text{ atm}}{760 \text{ torr}} = 3.36 \times 10^{-3}$$
 mol of methanol is needed to ensure that there is a liquid and gas phase. This equals 1.08 grams.

7.10. Since, according to equation 7.11, $\alpha_i = \frac{P_i}{P_{i^*}}$, and since $P_{i^*} = 760 \text{ torr}$, we can calculate α_i :

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AW Rasmussen

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1. Understanding the eBook Ball Solution Manual
 - The Rise of Digital Reading Ball Solution Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Ball Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
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 - Popular eBook Platforms
 - Features to Look for in an Ball Solution Manual
 - User-Friendly Interface
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 - Personalized Recommendations
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- Highlighting and Note-Taking Ball Solution Manual
- Interactive Elements Ball Solution Manual

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- Participating in Virtual Book Clubs
- Following Authors and Publishers Ball Solution Manual

9. Balancing eBooks and Physical Books Ball Solution Manual

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Ball Solution Manual

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Ball Solution Manual

- Setting Reading Goals Ball Solution Manual
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Ball Solution Manual

- Fact-Checking eBook Content of Ball Solution Manual
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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