

## Chapter 3 States of Matter

**Section 3.1 Solids, Liquids, and Gases****(pages 68–73)**

This section explains how materials are classified as solids, liquids, or gases. It also describes the behavior of these three states of matter.

**Reading Strategy (page 68)**

**Comparing and Contrasting** As you read about the states of matter, replace each letter in the diagram below with one of these phrases: definite volume, definite shape, variable volume, or variable shape. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

**Describing the States of Matter (pages 68–73)**

- What are the three common states of matter?  
a. solids      b. liquids      c. gases
- Is the following sentence true or false? The fact that a copper wire can be bent shows that some solids do not have a definite shape.  
True
- Circle the letter of each phrase that describes how particles at the atomic level are arranged within most solids.  
a. randomly arranged      ☒ b. packed close together  
☒ c. arranged in a regular pattern      d. spread far apart
- Is the following sentence true or false? A liquid takes the shape of its container. True
- What is the state of matter in which a material has neither a definite shape nor a definite volume? gas
- Compare and contrast the arrangement of particles at the atomic level for a liquid and a solid. Particles in a solid are packed close together in an orderly arrangement. The arrangement of particles in a liquid is more random.
- What determines the shape and volume of a gas? A gas takes the shape and volume of its container.
- On the sun, where temperatures are extremely high, matter exists in a state known as plasma.

## Section 3 1 Solids Liquids And Gases

**William Garnett**



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