

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a solution and the pH of a solution based on concentration.
2. Calculate pH and concentration of acids and bases and the pH of a solution based on concentration.
3. Calculate the pH of a solution (25°C).

Vocabulary

- | | | |
|-------------------|---|---------------------|
| 1. autoionization | 2. hydronium ion (H_3O^+) | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. neutral solution |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|---|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}_3\text{O}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}_3\text{O}^+]$ |

Part A Completion

Write the equation for the autoionization of water. What is the concentration of H^+ in pure water? What is the concentration of OH^- in pure water?

What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ H^+ ?

What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ OH^- ?

What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ H^+ ?

What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ OH^- ?

What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ H^+ ?

What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ OH^- ?

What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ H^+ ?

What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ OH^- ?

What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ H^+ ?

What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ OH^- ?

What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ H^+ ?

What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ OH^- ?

1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$

2. $1.0 \times 10^{-7} \text{ M}$

3. $1.0 \times 10^{-7} \text{ M}$

4. $1.0 \times 10^{-7} \text{ M}$

5. $1.0 \times 10^{-7} \text{ M}$

6. $1.0 \times 10^{-7} \text{ M}$

7. $1.0 \times 10^{-7} \text{ M}$

8. $1.0 \times 10^{-7} \text{ M}$

9. $1.0 \times 10^{-7} \text{ M}$

10. $1.0 \times 10^{-7} \text{ M}$

11. $1.0 \times 10^{-7} \text{ M}$

12. $1.0 \times 10^{-7} \text{ M}$

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