

## AP CHEMISTRY FINAL EXAM REVIEW 2

1. Is energy absorbed or released in an endothermic reaction?: Absorbed.
2. What is the sign of delta H for an exothermic reaction?: Exothermic reactions release energy. The sign of delta H is negative.
3. When you touch a flask and it feels hot, is this an endothermic or exothermic process?: Exothermic; heat is being given off.
4. Define specific heat capacity.: The amount of heat required to raise 1g of a substance to 1 degrees Celsius.
5.  $q=mc(\Delta T)$  : What does "q" stand for?: Heat (kJ)
6.  $q=mc(\Delta T)$  : What does "c" stand for?: Specific heat capacity (J / g degreesC)
7. What is a hydrogen ion donor called?: An acid.
8. What is a hydrogen ion acceptor called?: A base.
9. Feels slippery; acid or base?: Base.
10. Tastes sour; acid or base?: Acid.
11. Tastes bitter; acid or base?: Base.
12.  $pH < 7$  : acid, base, or neutral?: Acid
13.  $pH > 7$  : acid, base, or neutral?: Base
14. Define neutralization.: When an acid reacts with a base to form water and a salt.  $pH = 7$ .
15. Soap; acid or base?: Base (tastes bitter)
16. Ammonia, acid or base?: Base (drain cleaners are bases).
17. Lemonade; acid or base?: Acid (tastes sour)
18. How many valence electrons are in an atom of chlorine?: 7
19. How many valence electrons does aluminum have?: 3
20. Define Reaction Rate: The decrease in concentration of reactants or the increase in concentration of products with time.
21. How does increasing temperature affect the rate of reaction?: Increasing temperature causes the kinetic energy to increase. Thus, molecules are moving faster and colliding more frequently. This causes the rate of reaction to increase.
22. If I decrease the concentration of reactant, what happens to the rate of reaction?: Decreasing concentration leads to less reactants available for collision. Therefore, reaction rate decreases.
23. How does surface area influence reaction rate?: By increasing the surface area of reactant molecules, we are increasing the space or point of contact that these molecules have available

# 2nd Semester Chemistry Final Exam Review Answers

**Terry C. Jones**

## 2nd Semester Chemistry Final Exam Review Answers:

**Artificial Intelligence in Education. Posters and Late Breaking Results, Workshops and Tutorials, Industry and Innovation Tracks, Practitioners, Doctoral Consortium and Blue Sky** Andrew M. Olney, Irene-Angelica

Chounta, Zitao Liu, Olga C. Santos, Ig Ibert Bittencourt, 2024-07-01 This volume constitutes poster papers and late breaking results presented during the 25th International Conference on Artificial Intelligence in Education AIED 2024 which took place in Recife Brazil during July 8-12 2024. The 18 full papers and 92 short papers were carefully reviewed and selected from 200 submissions. They are organized in topical sections as follows: Part One: Blue Sky, Industry, Innovation and

Practitioner, WideAIED and Late Breaking Results; Part Two: Late Breaking Results, Doctoral Consortium, Workshops and Tutorials. *Quarterly Review of Distance Education* Michael Simonson, Charles Schlosser, 2018-10-01 The Quarterly Review of Distance Education is a rigorously refereed journal publishing articles, research briefs, reviews, and editorials dealing with the theories, research, and practices of distance education. The Quarterly Review publishes articles that utilize various methodologies that permit generalizable results which help guide the practice of the field of distance education in the public and private sectors. The Quarterly Review publishes full length manuscripts as well as research briefs, editorials, reviews of programs and scholarly works and columns. The Quarterly Review defines distance education as institutionally based formal education in which the learning group is separated and interactive technologies are used to unite the learning group.

*Resources in Education*, 1990-05 *Essentials of Physical Chemistry* Don Shillady, 2011-07-27 At a time when U.S. high school students are producing low scores in mathematics and science on international examinations, a thorough grounding in physical chemistry should not be considered optional for science undergraduates. Based on the author's thirty years of teaching, *Essentials of Physical Chemistry* merges coverage of calculus with chemistry and molecular physics in a friendly yet thorough manner. Reflecting the latest ACS guidelines, the book can be used as a one or two semester course and includes special topics suitable for senior projects. The book begins with a math and physics review to ensure all students start on the same level and then discusses the basics of thermodynamics and kinetics with mathematics tuned to a level that stretches students' abilities. It then provides material for an optional second semester course that shows students how to apply their enhanced mathematical skills in a brief historical development of the quantum mechanics of molecules. Emphasizing spectroscopy, the text is built on a foundation of quantum chemistry and more mathematical detail and examples. It contains sample classroom tested exams to gauge how well students know how to use relevant formulas and to display successful understanding of key concepts. Coupling the development of mathematical skills with chemistry concepts encourages students to learn mathematical derivations. Mini biographies of famous scientists make the presentation more interesting from a people point of view. Stating the basic concepts of quantum chemistry in terms of analogies provides a pedagogically useful technique. Covering key topics such as the critical point of a van der Waals gas, the Michaelis-Menten equation and the

entropy of mixing this classroom tested text highlights applications across the range of chemistry forensic science pre medical science and chemical engineering In a presentation of fundamental topics held together by clearly established mathematical models the book supplies a quantitative discussion of the merged science of physical chemistry *Reports from the Classroom* Sarah H. Huyvaert,1995 Nursing School Entrance Exam Review William H. Brown, Frederick A. Bettelheim, Jerry March, 2004-04

**5 Steps to a 5 AP Chemistry, 2012-2013 Edition** Richard H. Langley, John Moore, 2011-06-10 A Perfect Plan for the Perfect Score We want you to succeed on your AP exam That's why we've created this 5 step plan to help you study more effectively use your preparation time wisely and get your best score This easy to follow guide offers you a complete review of your AP course strategies to give you the edge on test day and plenty of practice with AP style test questions You'll sharpen your subject knowledge strengthen your thinking skills and build your test taking confidence with Full length practice exams modeled on the real test All the terms and concepts you need to know to get your best score Your choice of three customized study schedules so you can pick the one that meets your needs The 5 Step Plan helps you get the most out of your study time Step 1 Set Up Your Study Program Step 2 Determine Your Readiness Step 3 Develop the Strategies Step 4 Review the Knowledge Step 5 Build Your Confidence Topics include Basics Reactions and Periodicity Stoichiometry Gases Thermodynamics Spectroscopy Light and Electrons Bonding Solids Liquids and Intermolecular Forces Solutions and Colligative Properties Kinetics Equilibrium Electrochemistry Nuclear Chemistry Organic Chemistry Experimental

**World of Chemistry** Melvin D. Joesten, 2004 Journal of the Pennsylvania Academy of Science, 2001

**5 Steps to a 5 AP Chemistry, 2014-2015 Edition** Richard H. Langley, John Moore, 2013-08-02 A PERFECT PLAN for the PERFECT SCORE STEP 1 Set up your study plan with three customized study schedules STEP 2 Determine your readiness with an AP style diagnostic exam STEP 3 Develop the strategies that will give you the edge on test day STEP 4 Review the terms and concepts you need to score high STEP 5 Build your confidence with full length practice exams

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Official Report of the Student Curriculum Evaluation Committee, 1968-1969 University of Virginia. College of Arts and Sciences. Student Curriculum Evaluation Committee, 1969 *Iterations, II* Russell Batt, John W. Moore, 1987 *El-Hi Textbooks in Print*, 1977 *Biology* Edward C. Gruber, 1963 *College Chemistry* Harold Goldwhite, John R. Spielman, 1984 An outline of the basic concepts of chemistry includes discussions of scientific notation atomic structure chemical bonding and the periodic table

**Guide to Medical**

**and Dental Schools** Saul Wischnitzer,Edith Wischnitzer,2006 Updated with the latest facts and figures this school directory and guidance manual presents profiles of all accredited medical dental and osteopathic schools in the United States and Canada Here is information on tuition and fees application procedures and much more The book also includes the latest available full length model Medical College Admission Test MCAT with answers and selected questions with answers from recent Dental College Admission Tests DAT The authors offer advice on when and where to apply to dental and medical schools information on assessing one's chances for acceptance at the schools and all important test taking tips for students preparing to take the MCAT or DAT This new edition includes sample essays for medical school applications and a special section dedicated to high school students who are making long term plans for a medical or dental career *Multicultural applications of mastery learning* Güzver Yıldırın,2006 **Chemical Engineering Education** ,1989 *Journal of Developmental Education* ,2002

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