

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa.
2. Calculate the pH of weak acids and bases, and vice versa.
3. Calculate the percent ionization.

Vocabulary

- acid dissociation constant
- base dissociation constant

- conjugate acid-base pair
- ionization constant
- percent ionization

- acid ionization
- base ionization

Key Equations

$$K_a = \frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]}$$

(see Table 19.2)

$$K_b = \frac{[\text{B}][\text{OH}^-]}{[\text{BOH}]}$$

(see Table 19.2)

Part A Completion

Use the appropriate words to complete the following statements. Each blank should be filled with either the chemical name or formula, or a value.

1. The K_a of acetic acid is 1.8×10^{-5} .
2. The K_b of ammonia is 1.8×10^{-5} .
3. The K_a of acetic acid is 1.8×10^{-5} .
4. The K_b of ammonia is 1.8×10^{-5} .
5. The K_a of acetic acid is 1.8×10^{-5} .
6. The K_b of ammonia is 1.8×10^{-5} .
7. The K_a of acetic acid is 1.8×10^{-5} .
8. The K_b of ammonia is 1.8×10^{-5} .
9. The K_a of acetic acid is 1.8×10^{-5} .
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17. The K_a of acetic acid is 1.8×10^{-5} .
18. The K_b of ammonia is 1.8×10^{-5} .
19. The K_a of acetic acid is 1.8×10^{-5} .
20. The K_b of ammonia is 1.8×10^{-5} .

21. The K_a of acetic acid is 1.8×10^{-5} .
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38. The K_b of ammonia is 1.8×10^{-5} .
39. The K_a of acetic acid is 1.8×10^{-5} .
40. The K_b of ammonia is 1.8×10^{-5} .

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Wolfgang Guggemos



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