

## 19.5

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from  $\text{pH}$  and  $\text{pOH}$  and vice versa from  $\text{pH}$  and  $\text{pOH}$  given.
3. Calculate the  $\text{pH}$  of a weak acid.

## Vocabulary

- |                   |                               |                  |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. $\text{pH}$                | 6. $\text{pOH}$  |

## Key Equation

- |                                    |                                      |
|------------------------------------|--------------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pOH} = -\log[\text{OH}^-]$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} + \text{pOH} = 14$     |

## Part A Completion

Write the hydrogen ion concentration or autoionization equilibrium of the acids and bases that are indicated in the spaces provided. Be sure to include the correct state symbols.

1. What substance is \_\_\_\_\_, a weak hydrogen ion concentration of  $1.0 \times 10^{-4} \text{ M}$  is the concentration of this acidic substance in a 0.10 M solution? \_\_\_\_\_
2. Hydrochloric acid has a  $\text{pH}$  of \_\_\_\_\_, so it is \_\_\_\_\_ acidic. \_\_\_\_\_
3. \_\_\_\_\_, a concentration of \_\_\_\_\_, is a \_\_\_\_\_ acidic. \_\_\_\_\_
4. \_\_\_\_\_, a concentration of \_\_\_\_\_, is a \_\_\_\_\_ acidic. \_\_\_\_\_
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**Effiong Eyo**



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