

Multiple Choice:

Selecting the alternative that constitutes the statement or answers the question. What is most commonly observed with water?

1. How many moles of hydrogen gas are in a cylinder at typical conditions, 1000 L?
 a. 1
 b. 10
 c. 100
 d. 1000

2. At temperature T_1 , 1 mol of H_2O takes twice as long to evaporate as at temperature T_2 .
 a. $T_1 = 2T_2$
 b. $T_1 = 2T_2$
 c. $T_1 = 2T_2$
 d. $T_1 = 2T_2$

3. The equilibrium shift caused by the removal of:
 a. H_2O at low pressure
 b. H_2O at high pressure
 c. H_2O at low pressure
 d. H_2O at high pressure

4. For temperature T_1 , 1 mol of H_2O takes twice as long to evaporate as at temperature T_2 .
 a. $T_1 = 2T_2$
 b. $T_1 = 2T_2$
 c. $T_1 = 2T_2$
 d. $T_1 = 2T_2$

5. What is the state of the water in the following system? H_2O is in the liquid state.
 a. liquid water
 b. solid water
 c. gaseous water
 d. solid water

6. What is the boiling point of water?
 a. 100°C
 b. 100°C
 c. 100°C
 d. 100°C

7. A sample of water is heated to 100°C at the boiling point. At this point, the water is in the liquid state.
 a. 100°C
 b. 100°C
 c. 100°C
 d. 100°C

8. For temperature T_1 , 1 mol of H_2O takes twice as long to evaporate as at temperature T_2 .
 a. $T_1 = 2T_2$
 b. $T_1 = 2T_2$
 c. $T_1 = 2T_2$
 d. $T_1 = 2T_2$

9. At temperature T_1 , 1 mol of H_2O takes twice as long to evaporate as at temperature T_2 .
 a. $T_1 = 2T_2$
 b. $T_1 = 2T_2$
 c. $T_1 = 2T_2$
 d. $T_1 = 2T_2$

10. The equilibrium constant for the reaction $\text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{O}$ is 1.0 .
 a. 1.0
 b. 1.0
 c. 1.0
 d. 1.0

11. The equilibrium constant for the reaction $\text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{O}$ is 1.0 .
 a. 1.0
 b. 1.0
 c. 1.0
 d. 1.0

Multiple Choice Stoichiometry Test

Sasha Alcott



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