

# Chemistry Study Guide

# Periodic Table of the Elements

<b>H</b> 1 Hydrogen																	<b>He</b> 2 Helium
<b>Li</b> 3 Lithium	<b>Be</b> 4 Beryllium											<b>B</b> 5 Boron	<b>C</b> 6 Carbon	<b>N</b> 7 Nitrogen	<b>O</b> 8 Oxygen	<b>F</b> 9 Fluorine	<b>Ne</b> 10 Neon
<b>Na</b> 11 Sodium	<b>Mg</b> 12 Magnesium											<b>Al</b> 13 Aluminum	<b>Si</b> 14 Silicon	<b>P</b> 15 Phosphorus	<b>S</b> 16 Sulfur	<b>Cl</b> 17 Chlorine	<b>Ar</b> 18 Argon
<b>K</b> 19 Potassium	<b>Ca</b> 20 Calcium	<b>Sc</b> 21 Scandium	<b>Ti</b> 22 Titanium	<b>V</b> 23 Vanadium	<b>Cr</b> 24 Chromium	<b>Mn</b> 25 Manganese	<b>Fe</b> 26 Iron	<b>Co</b> 27 Cobalt	<b>Ni</b> 28 Nickel	<b>Cu</b> 29 Copper	<b>Zn</b> 30 Zinc	<b>Ga</b> 31 Gallium	<b>Ge</b> 32 Germanium	<b>As</b> 33 Arsenic	<b>Se</b> 34 Selenium	<b>Br</b> 35 Bromine	<b>Kr</b> 36 Krypton
<b>Rb</b> 37 Rubidium	<b>Sr</b> 38 Strontium	<b>Y</b> 39 Yttrium	<b>Zr</b> 40 Zirconium	<b>Nb</b> 41 Niobium	<b>Mo</b> 42 Molybdenum	<b>Tc</b> 43 Technetium	<b>Ru</b> 44 Ruthenium	<b>Rh</b> 45 Rhodium	<b>Pd</b> 46 Palladium	<b>Ag</b> 47 Silver	<b>Cd</b> 48 Cadmium	<b>In</b> 49 Indium	<b>Sn</b> 50 Tin	<b>Sb</b> 51 Antimony	<b>Te</b> 52 Tellurium	<b>I</b> 53 Iodine	<b>Xe</b> 54 Xenon
<b>Cs</b> 55 Cesium	<b>Ba</b> 56 Barium		<b>Hf</b> 72 Hafnium	<b>Ta</b> 73 Tantalum	<b>W</b> 74 Tungsten	<b>Re</b> 75 Rhenium	<b>Os</b> 76 Osmium	<b>Ir</b> 77 Iridium	<b>Pt</b> 78 Platinum	<b>Au</b> 79 Gold	<b>Hg</b> 80 Mercury	<b>Tl</b> 81 Thallium	<b>Pb</b> 82 Lead	<b>Bi</b> 83 Bismuth	<b>Po</b> 84 Polonium	<b>At</b> 85 Astatine	<b>Rn</b> 86 Radon
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		<div> <div> Actinium 89 Ac </div> <div> Thorium 90 Th </div> <div> Protactinium 91 Pa </div> <div> Uranium 92 U </div> <div> Neptunium 93 Np </div> <div> Plutonium 94 Pu </div> <div> Americium 95 Am </div> <div> Curium 96 Cm </div> <div> Berkelium 97 Bk </div> <div> Californium 98 Cf </div> <div> Einsteinium 99 Es </div> <div> Fermium 100 Fm </div> <div> Mendelevium 101 Md </div> <div> Nobelium 102 No </div> <div> Lawrencium 103 Lr </div> </div>															

## Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles 🎯

## Matter:

Anything that has **mass** and **occupies space** ✖️🌐

Examples: air, gases, liquids, solids!

## Atoms:

Atoms are the **smallest, indivisible particle** 🧪👉

- **Outer part** → *shell/orbit/energy level* → carries the electrons ( $e^-$ )
- **Inner part** → *nucleus* → houses the protons ( $p^+$ ) and neutrons ( $n$ )
- Outer & Inner part → together these are the **subatomic particles**

## Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	$p^+$	1	+1	Nucleus
Neutron	$n$	1	0	Nucleus
Electron	$e^-$	$\sim 0$	-1	Orbit/Shell

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