

Acids & Bases HOMEWORK

Homework

1. Propose the correct ion search of the strong acid listed below. Propose the incorrect ion search of the strong base listed below.

- a. **HF**: **Hydrofluoric acid**
- b. **HI**: **Hydroiodic acid**
- c. **HNO₃**: **nitric acid**
- d. **HClO₄**: **perchloric acid**
- e. **HBrO₃**: **hypobromous acid**
- f. **HClO₄**: **chloric acid**
- g. **H₂SO₄**: **hydrochloric acid**
- h. **H₂SO₄**: **hydrobromic acid**
- i. **H₂SO₄**: **hydroiodic acid**
- j. **H₂SO₄**: **hydrochloric acid**

2. Of the acids and bases listed below, write the ones that completely dissociate in water.

Acid	Base	Mixture	Neutral	Acidic	Basic	Neutral	Basic
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3. Write the dissociation for **Na₃PO₄**.

4. **NaClO₄** will react with water and lose a ClO_4^- ion in the process.

a. Write a balanced equation for this reaction: $\text{NaClO}_4 + \text{H}_2\text{O} \longrightarrow \text{Na}^+ + \text{HClO}_4^-$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating HClO_4^- in a weak acid.

c. Dissolve **NaClO₄** in a portion of water and a portion of ethanol? **commonly** dissociating in both at $\text{pH} 7$?

d. In your **reaction** prediction, is **NaClO₄** acting as a **Brønsted-Lowry** acid or a **Brønsted-Lowry** base? **Brønsted-Lowry acid**. Because it is a portion of a strong acid.

e. Identify the conjugate base of **NaClO₄**: **ClO₄⁻**

f. Identify **NaClO₄** as **strong**, **weak**, **intermediate**, or **none**: **strong**, because it is a portion of a strong acid. **NaClO₄**

5. **Li⁺** will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{Li}^+ + \text{H}_2\text{O} \longrightarrow \text{H}^+ + \text{LiOH}$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating LiOH in a weak base.

c. Dissolve **LiOH** in a portion of water and a portion of ethanol? **commonly** dissociating in both at $\text{pH} 7$?

d. In your **reaction** prediction, is **LiOH** acting as a **Brønsted-Lowry** acid or a **Brønsted-Lowry** base? **Brønsted-Lowry base**. Because it is a portion of a strong base.

e. Identify the conjugate acid of **LiOH**: **Li⁺**

f. Identify **LiOH** as **strong**, **weak**, **intermediate**, or **none**: **strong**

6. **Na₃PO₄** will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{Na}_3\text{PO}_4 + \text{H}_2\text{O} \longrightarrow \text{Na}^+ + \text{H}_3\text{PO}_4^-$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating H_3PO_4^- in a weak acid.

c. Dissolve **Na₃PO₄** in a portion of water and a portion of ethanol? **commonly** dissociating in both at $\text{pH} 7$?

d. In your **reaction** prediction, is **Na₃PO₄** acting as a **Brønsted-Lowry** acid or a **Brønsted-Lowry** base? **Brønsted-Lowry acid**. Because it is a portion of a strong acid.

e. Identify the conjugate base of **Na₃PO₄**: **PO₄³⁻**

f. Identify **Na₃PO₄** as **strong**, **weak**, **intermediate**, or **none**: **strong**, because it is a portion of a strong acid. **Na₃PO₄** is like **NaClO₄**.

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Society of Chemical Industry (Great Britain)

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Table of Contents 19 3 Acids And Bases Answer Key Chemistry

1. Understanding the eBook 19 3 Acids And Bases Answer Key Chemistry
 - The Rise of Digital Reading 19 3 Acids And Bases Answer Key Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 3 Acids And Bases Answer Key Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 3 Acids And Bases Answer Key Chemistry
 - User-Friendly Interface

4. Exploring eBook Recommendations from 19 3 Acids And Bases Answer Key Chemistry
 - Personalized Recommendations
 - 19 3 Acids And Bases Answer Key Chemistry User Reviews and Ratings
 - 19 3 Acids And Bases Answer Key Chemistry and Bestseller Lists
5. Accessing 19 3 Acids And Bases Answer Key Chemistry Free and Paid eBooks
 - 19 3 Acids And Bases Answer Key Chemistry Public Domain eBooks
 - 19 3 Acids And Bases Answer Key Chemistry eBook Subscription Services
 - 19 3 Acids And Bases Answer Key Chemistry Budget-Friendly Options
6. Navigating 19 3 Acids And Bases Answer Key Chemistry eBook Formats
 - ePUB, PDF, MOBI, and More
 - 19 3 Acids And Bases Answer Key Chemistry Compatibility with Devices
 - 19 3 Acids And Bases Answer Key Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 3 Acids And Bases Answer Key Chemistry
 - Highlighting and Note-Taking 19 3 Acids And Bases Answer Key Chemistry
 - Interactive Elements 19 3 Acids And Bases Answer Key Chemistry
8. Staying Engaged with 19 3 Acids And Bases Answer Key Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 3 Acids And Bases Answer Key Chemistry
9. Balancing eBooks and Physical Books 19 3 Acids And Bases Answer Key Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 3 Acids And Bases Answer Key Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 19 3 Acids And Bases Answer Key Chemistry
 - Setting Reading Goals 19 3 Acids And Bases Answer Key Chemistry
 - Carving Out Dedicated Reading Time

- 12. Sourcing Reliable Information of 19 3 Acids And Bases Answer Key Chemistry
 - Fact-Checking eBook Content of 19 3 Acids And Bases Answer Key Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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