

## **Carbonate Ores Method**

### **Hypothesis:**

"The amount of carbonate in an ore sample can be measured by adding hydrochloric acid to the sample until it stops fizzing."

### **Equipment:**

- 100ml beaker
- Samples of different carbonate ores labeled 10%, 15%, 20%, 25%, 30% of copper carbonate.
- Measuring cylinder
- Stopwatch
- 2M Hydrochloric acid
- Spatula

### **Method:**

1. Put two spatulas of an ore sample into the beaker.
2. Measure 25ml of 2M hydrochloric acid using the measuring cylinder.
3. Pour the dilute hydrochloric acid into the beaker and start the stopwatch. Swirl the beaker once.
4. Measure the time taken for the fizzing to stop.
5. Repeat for the other samples.

### **Variables:**

**Independent Variable:** Percentage of Copper Carbonate in the ore (%)

**Dependent Variable:** Time taken for the fizzing to stop (s)

**Control Variables:** Volume of acid, concentration of acid.

### **Risk Assessment:**

**Hazard:** Hydrochloric acid

**Risk:** Could get into your eyes and cause irritation.

**Control:** Wear safety glasses at all times when doing the investigation

**Hazard:** Copper Carbonate

**Risk:** It is a harmful chemical and could get into your eyes or on your skin.

**Control:** Wear safety glasses at all times when doing the investigation. If you get it on your skin, wash off with cold water.

### **Title**

Carbonate Ores Method

### **Web address**

<http://www.tes.co.uk/teaching-resource>

# Method Of Carbonate Ores Isa Practical Aqa

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**Chemical News and Journal of Industrial Science** ,1877      **The Chemical News and Journal of Industrial Science** William Crookes,James H. Gardiner,Gerald Druce,H. W. Blood-Ryan,1877      **The Chemical News** ,1877      **The Mining Magazine** ,1927      **Beeton's Dictionary of universal information; comprising a complete summary of the moral, mathematical, physical and natural sciences [&c., ed. by S.O. Beeton and J. Sherer. Wanting pt. 13].**  
Samuel Orchart Beeton,1870      **Corpus Juris** William Mack,William Benjamin Hale,1919

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