

Chemistry 2nd Semester Exam Review Sheet:**Chapters 9, 10, 11, 12, 13, 14, 15, and 16****Chapter 9: Chemical Quantities**

1. Solve stoichiometry problems using grams \rightarrow moles \rightarrow moles \rightarrow grams calculations. Examples below:
a. Calculate moles/grams of $(\text{NH}_4)_2\text{SO}_4$ produced when 30.0 moles of NH_3 react with H_2SO_4 .



$$30.0 \text{ moles} \times 1 \text{ mol/2 mol} = 15.0 \text{ moles}$$

$$15.0 \text{ moles } (\text{NH}_4)_2\text{SO}_4 \times 132.2 \text{ g/mol/mole} = 1983 = 1980 \text{ grams (to 3 s.f.)}$$

b. $\text{Na}_2\text{CO}_3 + 2\text{HNO}_3 \rightarrow 2\text{NaNO}_3 + \text{CO}_2 + \text{H}_2\text{O}$. How many moles/grams of sodium nitrate are produced by the complete reaction of 100. grams of sodium carbonate? Moles of nitric acid needed to produce 500.0 grams of sodium nitrate? *Be able to calculate grams or moles of any reactant or product given a quantity of reactant or product.*

Grams \rightarrow Moles \rightarrow Moles \rightarrow Grams

$$100. \text{ g } \text{Na}_2\text{CO}_3 \times 1 \text{ mol/106.0 g} \times 2 \text{ mol } \text{NaNO}_3 / 1 \text{ mol } \text{Na}_2\text{CO}_3 \times 85.00 \text{ g/mol} = 160.4 \text{ g}$$

$$500.0 \text{ g } \text{NaNO}_3 \times 1 \text{ mol/85.00 g} \times 2 \text{ mol / 2 mol} \times 63.02 \text{ g/mol} = 370.7 \text{ grams}$$

2. Calculate percent yield of products. Example below using (b) above.

What is the theoretical yield of NaNO_3 if 100.0 grams of Na_2CO_3 are used in the reaction? What is the percent yield of NaNO_3 if 130.5 grams of NaNO_3 are actually collected?

Theoretical yield = 160.4 grams (answer from above)

$$\text{Percent Yield} = (\text{Actual Yield/Theoretical Yield}) \times 100 = (130.5 \text{ g}/160.4 \text{ g}) \times 100 = 81.36\%$$

3. What is a limiting reactant in a chemical reaction? The reactant that is used up first, thereby limiting the amount of product that can be formed.

Chapter 10: Modern Atomic Theory

4. Describe Rutherford's model of the atom. What could it not explain?

Nucleus (very small) consisting of protons with electrons in orbits around the nucleus. Could not explain how the electrons remained in orbit.

5. Define electromagnetic radiation, wavelength, and frequency.

6. How do atoms gain energy to become excited? Absorb a photon of energy.

How do they lose energy to return to their ground state? Release a photon of same energy as that absorbed.

7. What is the evidence that energy absorbed by the hydrogen atom is quantized? Specific wavelengths of light seen in emission spectra for atoms (page 285-86)

8. Describe the Bohr model of the atom. How did it improve on Rutherford's model?

Same as Rutherford with respect to nucleus. Electrons located on orbits at specific distances from nucleus.

9. Draw the orbital diagram (with the arrows) for an element: for example, phosphorus, sodium, and fluorine.

See book, notes, probs 54, 96

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E. Oñate

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REA's Authoritative Guide to Graduate Schools Research and Education Association, Rea, Staff of Research Education Association, 1998-01-01 REAs reference book profiles top graduate schools in over sixty fields of study including engineering, biology, psychology and chemistry. The profiles have clear, easy to read comparison charts that give details to help you select the best graduate school for you. Contains information on enrollment, admissions requirements, financial aid, tuition and much more. This book is a helpful guide to students who are considering graduate school.

Source Book for Chemistry

Teachers William Thomas Lippincott, 1981 **Resources in education**, 1987-09 **The American Secondary School**

Leslie Owen Taylor, 1960 Engineering Education, 1983 **Cornell University Courses of Study** Cornell University, 2003 University of Colorado at Denver Catalog University of Colorado at Denver, 2009 **Summer School**

Number Southwestern at Memphis, 1933 **Catalog Issue** University of Colorado, 1974 THE Journal, 1984 Current Index to Journals in Education, 1986 Proceedings, 1984 **The School Review**, 1914 Science Books & Films, 1980

American Journal of Physics, 1994 **Computer Aided Training in Science and Technology** E. Oñate, 1990

Computer-assisted Instruction Albert E. Hickey, 1968 Abridged Index Medicus, 1976 Self-study Report: Self-study report : appendices -v. 3. Abstract of the Self-study report University of Illinois at Urbana-Champaign, 1979 **Paperbound Books in Print**, 1966

This book delves into 2nd Semester Chemistry Exam Review Fot Plato. 2nd Semester Chemistry Exam Review Fot Plato is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into 2nd Semester Chemistry Exam Review Fot Plato, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to 2nd Semester Chemistry Exam Review Fot Plato
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- Chapter 3: 2nd Semester Chemistry Exam Review Fot Plato in Everyday Life
- Chapter 4: 2nd Semester Chemistry Exam Review Fot Plato in Specific Contexts
- Chapter 5: Conclusion

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6. In chapter 5, the author will draw a conclusion about 2nd Semester Chemistry Exam Review Fot Plato. This chapter will summarize the key points that have been discussed throughout the book.
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there are three states of matter solid liquid and gas in chapter 1 of ncert solutions for class 12 chemistry solid characteristics are studied solids are defined by their mass shape and volume intermolecular distance and force are both short yet intermolecular force is strong

