



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

Key Vocabulary

<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base
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Key Equations

- $\text{pH} = -\log[\text{H}^+]$
- $\text{pH} = \log \frac{1}{[\text{H}^+]}$

$$\text{pH} = -\log[\text{H}^+]$$

Practice Problems

1. Calculate the pH of a 0.001 M HCl solution.

Answers

1. 3.00	2. 1.00	3. 1.00
4. 1.00	5. 1.00	6. 1.00
7. 1.00	8. 1.00	9. 1.00
10. 1.00	11. 1.00	12. 1.00
13. 1.00	14. 1.00	15. 1.00
16. 1.00	17. 1.00	18. 1.00
19. 1.00	20. 1.00	21. 1.00
22. 1.00	23. 1.00	24. 1.00
25. 1.00	26. 1.00	27. 1.00
28. 1.00	29. 1.00	30. 1.00
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