

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from pH and pOH , and vice versa from pH and pOH .
3. Calculate the pH of a weak acid.

Vocabulary

- | | | |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. pH | 6. pOH |

Key Equation

- | | |
|------------------------------------|----------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = 14 - \text{pOH}$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = 14 - \text{pOH}$ |

Part A Completion

Write the hydrogen ion concentration or autoionization constant, if the missing quantity that is indicated in the partial equation below represents either an autoionization constant.

1. Water autoionizes to H^+ and hydroxide ions. The autoionization constant K_w of the autoionization of water which produces an H^+ ion and a hydroxide ion is 1.0×10^{-14} .
2. Hydrochloric acid dissociates from HCl to H^+ and Cl^- . The K_a of HCl is 1.0×10^{-7} . The concentration of H^+ in the solution is 1.0×10^{-7} M. The pH of the solution is 7 .
3. The pH of a solution is 4 . The concentration of H^+ in the solution is 1.0×10^{-4} M. The pOH of the solution is 10 .
4. The pH of a solution is 4 . The concentration of H^+ in the solution is 1.0×10^{-4} M. The pOH of the solution is 10 .
5. The pH of a solution is 4 . The concentration of H^+ in the solution is 1.0×10^{-4} M. The pOH of the solution is 10 .

6. $\text{pH} = 4$
7. $\text{pH} = 10$
8. $\text{pH} = 10$
9. $\text{pH} = 10$
10. $\text{pH} = 10$
11. $\text{pH} = 10$
12. $\text{pH} = 10$
13. $\text{pH} = 10$
14. $\text{pH} = 10$
15. $\text{pH} = 10$
16. $\text{pH} = 10$
17. $\text{pH} = 10$
18. $\text{pH} = 10$

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