

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- HCl ~~hydrochloric acid~~
- HBr ~~hydrobromic acid~~
- HNO₃ ~~nitric acid~~
- H₂SO₄ ~~sulfuric acid~~
- Strontium hydroxide ~~LiOH~~
- potassium hydroxide ~~KOH~~
- Sodium hydroxide ~~NaOH~~
- calcium hydroxide ~~Ca(OH)₂~~
- ammonium hydroxide ~~Li(OH)₄~~

~~HClO₄ is a STRONG acid & it does not completely dissociate in water.~~

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

HCl

HBr

HNO₃

Mg(OH)₂

HNO₂

NaOH

NaClO₄

H₂SO₄

3. What is the formula for hydronium? H₃O⁺

4. H₂C₂O₄ will react with water and lose a H⁺ ion in the process.

- Write a balanced equation for this reaction: $H_2C_2O_4 + H_2O \rightleftharpoons H_3O^+ + HC_2O_4^-$
- Would you expect a large or small equilibrium constant for this reaction? small, because H₂C₂O₄ is a weak acid.
- Would you define H₂C₂O₄ as a proton donor or a proton acceptor? proton donor, because it is losing a H⁺ ion.
- In your written equation, is H₂C₂O₄ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? Brønsted-Lowry acid, because it is a proton donor.
- Identify the conjugate base of H₂C₂O₄: HC₂O₄⁻
- Identify H₂C₂O₄ as polyprotic, amphoteric, or neither. polyprotic, because it contains multiple H⁺ atoms.

5. F⁻ will react with water and gain a H⁺ ion in the process.

- Write a balanced equation for this reaction: $F^- + H_2O \rightleftharpoons HF + OH^-$
- Would you expect a large or small equilibrium constant for this reaction? small, because F⁻ is a weak base.
- Would you define F⁻ as a proton donor or a proton acceptor? proton acceptor, because it is gaining a H⁺ ion.
- In your written equation, is F⁻ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? Brønsted-Lowry base, because it is a proton acceptor.
- Identify the conjugate acid of F⁻: HF
- Identify F⁻ as polyprotic, amphoteric, or neither: neither

6. HSO₄⁻ will react with water and lose a H⁺ ion in the process.

- Write a balanced equation for this reaction: $HSO_4^- + H_2O \rightleftharpoons H_3O^+ + SO_4^{2-}$
- Would you expect a large or small equilibrium constant for this reaction? small, because HSO₄⁻ is a weak acid.
- Would you define HSO₄⁻ as a proton donor or a proton acceptor? proton donor, because it is losing a H⁺ ion.
- In your written equation, is HSO₄⁻ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? Brønsted-Lowry acid, because it is a proton donor.
- Identify the conjugate base of HSO₄⁻: SO₄²⁻
- Identify HSO₄⁻ as polyprotic, amphoteric, or neither: amphoteric, because it has a negative charge (and can take a proton) and it has H⁺ atoms (and can give one away).

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