

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, given  $K_a$  or  $K_b$ .
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, given  $K_a$  or  $K_b$ .
3. Calculate the pH of weak acid and base solutions.

## Vocabulary

- acid dissociation constant
- base dissociation constant

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- acid dissociation constant

## Key Equations

- 1.  $K_a = \frac{[H^+][A^-]}{[HA]}$  (Equation 19.1)
- 2.  $K_b = \frac{[OH^-][BH^+]}{[B]}$  (Equation 19.2)

- 3.  $K_a = \frac{[H^+][A^-]}{[HA]}$  (Equation 19.1)
- 4.  $K_b = \frac{[OH^-][BH^+]}{[B]}$  (Equation 19.2)

## Part A Completion

Write the chemical formula for the acid or base corresponding to the  $K_a$  or  $K_b$  value given. Write the chemical formula for the acid or base corresponding to the  $K_a$  or  $K_b$  value given.

1.  $K_a = 1.8 \times 10^{-4}$  (Equation 19.1)

2.  $K_b = 1.8 \times 10^{-4}$  (Equation 19.2)

3.  $K_a = 1.8 \times 10^{-4}$  (Equation 19.1)

4.  $K_b = 1.8 \times 10^{-4}$  (Equation 19.2)

5.  $K_a = 1.8 \times 10^{-4}$  (Equation 19.1)

6.  $K_b = 1.8 \times 10^{-4}$  (Equation 19.2)

7.  $K_a = 1.8 \times 10^{-4}$  (Equation 19.1)

8.  $K_b = 1.8 \times 10^{-4}$  (Equation 19.2)

9.  $K_a = 1.8 \times 10^{-4}$  (Equation 19.1)

10.  $K_b = 1.8 \times 10^{-4}$  (Equation 19.2)

11.  $K_a = 1.8 \times 10^{-4}$  (Equation 19.1)

12.  $K_b = 1.8 \times 10^{-4}$  (Equation 19.2)

13.  $K_a = 1.8 \times 10^{-4}$  (Equation 19.1)

14.  $K_b = 1.8 \times 10^{-4}$  (Equation 19.2)

15.  $K_a = 1.8 \times 10^{-4}$  (Equation 19.1)

16.  $K_b = 1.8 \times 10^{-4}$  (Equation 19.2)

17.  $K_a = 1.8 \times 10^{-4}$  (Equation 19.1)

18.  $K_b = 1.8 \times 10^{-4}$  (Equation 19.2)

19.  $K_a = 1.8 \times 10^{-4}$  (Equation 19.1)

20.  $K_b = 1.8 \times 10^{-4}$  (Equation 19.2)

21.  $K_a = 1.8 \times 10^{-4}$  (Equation 19.1)

22.  $K_b = 1.8 \times 10^{-4}$  (Equation 19.2)

# 19 2 Section Review Hydrogen Ions And Acidity

**Patricia M. Irving**



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