

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode  $T \gg \theta_m$ ) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute  $\frac{3}{2}R$  to the molar heat capacity. If we denote the number of active rotational modes by  $v_f^*$  (so for most molecules at normal temperatures  $v_f^* = 2$  for linear molecules, and 3 for non-linear molecules), then

# Physical Chemistry 9th Edition Instructors Solution Guide

**Thomas Griffiths**

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*Children's Books in Print, 2007, 2006* *Instructor's Resource Manual for Woolfolk, Educational Psychology, Ninth Edition* Angela M. O'Donnell, 2004 *Subject Guide to Books in Print, 2001* *Educational Times, 1884* *Instructor's Resource Manual to Accompany The Little, Brown Handbook, Ninth Edition* Jane E. Aaron, Robert A. Schwegler, 1986

Forthcoming Books Rose Arny, 2003 *An Introduction to Physical Science* James T. Shipman, Jerry D. Wilson, Aaron W. Todd, 2002-07-08 This text presents a survey of the physical sciences physics chemistry astronomy meteorology and geology for non science majors. Topics are treated both descriptively and quantitatively providing flexibility for instructors who wish to emphasize a highly descriptive approach a highly quantitative approach or any spectrum in between.

**The English Catalogue of Books, 1899** *Instructor, 1972-08* *Elements of General and Biological Chemistry, Laboratory Manual* John R. Holum, 1994-10-19 This updated edition explains recent advances in environmental studies and in the molecular basis of life. Suitable for students interested in the health care field as well as those who want to know how nature and human life work at the molecular level the book begins by providing readers with a solid background in formulas structures equations solutions and equilibria. A number of topics are introduced early such as molarity and are discussed in more detail in later chapters. Each chapter contains a summary as well as review exercises.

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