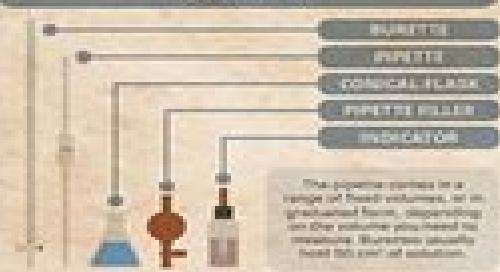


# CHEMISTRY TECHNIQUES: TITRATION

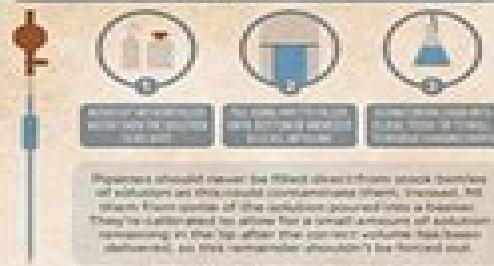
Used to determine the concentration of a particular solution, by measuring how much of a solution of known concentration reacts with a known volume of it.

## EQUIPMENT



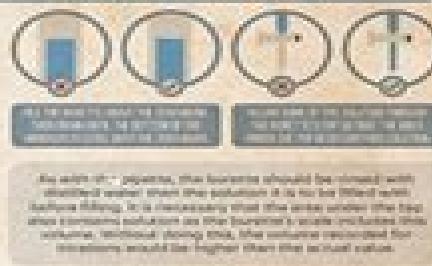
The burette contains a range of glassware, all in graduated form, depending on the volume you need to measure. Burettes often have two sets of markings.

## USING A PIPETTE



Pipettes should never be tilted away from the neck of the solution and should be rinsed with water. The volume of the solution is determined by the number of drops of liquid added. Pipettes often need to be rinsed to remove any solution remaining in the tip when they are being used. The volume measured should be recorded as the number of drops.

## FILLING THE BURETTE



As with a pipette, the burette should be rinsed with distilled water when the solution is no longer used. Before filling, it is necessary that the water used is free from dissolved carbon dioxide and therefore it is best to use distilled water. When measuring titrations, the burette should be higher than the actual value.

## USING INDICATORS



In acid-base titrations, a range of indicators can be used. These are solutions which change colour at a specific pH, and can be used to precisely identify when the neutralisation reaction is complete. There are different indicators available for different acid-base combinations.

## CARRYING OUT THE TITRATION



Carrying out the titration, one tip of the burette is exposed to allow the solution inside to flow into a known volume of the solution in the conical flask. The volume of solution from the burette required to reach the end point is recorded. A rough titration is usually followed by more accurate runs. Multiple titrations are carried out until consistent titres are obtained (within 0.50 cm<sup>3</sup> of each other).

## CARRYING OUT CALCULATIONS

$$n = CV$$

NO. OF MOLES CONCENTRATION VOLUME

1 Calculate number of moles of solution used from the burette.

2 Calculate the number of moles of solution in the conical flask using the equation for the reaction and reading value.

3 Calculate the concentration of the solution in the conical flask by rearranging the equation to  $C = n / V$ .



# Techniques Of Chemistry

**Gilbert J. Sloan, Andrew R. McGhie**

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