

Multiple Choice:

Selecting the alternative that constitutes the statement or answers the question. What is most commonly observed with water?

1. How many moles of hydroxide groups are a molecule of typical silicate tetrahedron?
 a. 1
 b. 2
 c. 3
 d. 4

2. An element X_2 ($100g$) = $100g$, what factor indicates molecular composition?
 a. 200
 b. 100
 c. 50
 d. 25

3. The unpaired electron in the ferrous ion:
 a. Spin up state
 b. Spin down state
 c. Spin up/down
 d. Spin up/down

4. Full element X_2 ($100g$) = $100g$, how many moles of X_2 are present in the sample?
 a. 100 moles
 b. 50 moles
 c. 25 moles
 d. 100 moles

5. What is the study of the relationship between chemical and physical properties of materials?
 a. Chemical composition
 b. Chemical structure
 c. Physical properties
 d. Chemical structure

6. What is the basic unit of matter?
 a. 100 moles
 b. 100 moles
 c. 100 moles
 d. 100 moles

7. A compound's empirical formula is CH_2 , at the same time, it is the same as the molecular formula.
 a. 100
 b. 100
 c. 100
 d. 100

8. Full element X_2 ($100g$) = $100g$, how many moles of X_2 are present in the sample?
 a. 100 moles
 b. 100 moles
 c. 100 moles
 d. 100 moles

9. An element X_2 ($100g$) = $100g$, what factor indicates molecular composition?
 a. 100
 b. 100
 c. 100
 d. 100

10. The molecular formula of a compound is CH_2 , at the same time, it is the same as the molecular formula.
 a. 100
 b. 100
 c. 100
 d. 100

11. The molecular formula of a compound is CH_2 , at the same time, it is the same as the molecular formula.
 a. 100
 b. 100
 c. 100
 d. 100

Multiple Choice Stoichiometry Test

John T. Moore, Richard H. Langley

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located on the right side of the slide.

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