

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and  $pK_a$  and  $pK_b$  for weak acids and bases.
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and  $pK_a$  and  $pK_b$  for weak acids and bases.
3. Calculate the  $pH$  of weak acid and base solutions.

## Vocabulary

- |                 |                                      |                 |
|-----------------|--------------------------------------|-----------------|
| acid            | acid dissociation constant ( $K_a$ ) | acid ionization |
| acidic solution | acidic solution                      | acidic solution |
| acidic solution | acidic solution                      | acidic solution |

## Key Equations

- |   |  |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

## Part A Completion

Write the  $K_a$  expression for the acid dissociation constant for each of the following acids. Write the  $K_b$  expression for the base dissociation constant for each of the following bases.

1.  $HNO_3$  (strong acid)
2.  $H_2SO_4$  (strong acid)
3.  $HClO_4$  (strong acid)
4.  $H_2CO_3$  (weak acid)
5.  $H_2O$  (amphoteric)
6.  $H_2O$  (amphoteric)
7.  $H_2O$  (amphoteric)
8.  $H_2O$  (amphoteric)
9.  $H_2O$  (amphoteric)
10.  $H_2O$  (amphoteric)
11.  $H_2O$  (amphoteric)
12.  $H_2O$  (amphoteric)
13.  $H_2O$  (amphoteric)
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16.  $H_2O$  (amphoteric)
17.  $H_2O$  (amphoteric)
18.  $H_2O$  (amphoteric)
19.  $H_2O$  (amphoteric)
20.  $H_2O$  (amphoteric)

1.  $HNO_3$
2.  $H_2SO_4$
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20.  $H_2O$

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