

Chemistry Study Guide

Periodic Table of the Elements

H 1 Hydrogen																	He 2 Helium
Li 3 Lithium	Be 4 Beryllium											B 5 Boron	C 6 Carbon	N 7 Nitrogen	O 8 Oxygen	F 9 Fluorine	Ne 10 Neon
Na 11 Sodium	Mg 12 Magnesium											Al 13 Aluminum	Si 14 Silicon	P 15 Phosphorus	S 16 Sulfur	Cl 17 Chlorine	Ar 18 Argon
K 19 Potassium	Ca 20 Calcium	Sc 21 Scandium	Ti 22 Titanium	V 23 Vanadium	Cr 24 Chromium	Mn 25 Manganese	Fe 26 Iron	Co 27 Cobalt	Ni 28 Nickel	Cu 29 Copper	Zn 30 Zinc	Ga 31 Gallium	Ge 32 Germanium	As 33 Arsenic	Se 34 Selenium	Br 35 Bromine	Kr 36 Krypton
Rb 37 Rubidium	Sr 38 Strontium	Y 39 Yttrium	Zr 40 Zirconium	Nb 41 Niobium	Mo 42 Molybdenum	Tc 43 Technetium	Ru 44 Ruthenium	Rh 45 Rhodium	Pd 46 Palladium	Ag 47 Silver	Cd 48 Cadmium	In 49 Indium	Sn 50 Tin	Sb 51 Antimony	Te 52 Tellurium	I 53 Iodine	Xe 54 Xenon
Cs 55 Cesium	Ba 56 Barium		Hf 72 Hafnium	Ta 73 Tantalum	W 74 Tungsten	Re 75 Rhenium	Os 76 Osmium	Ir 77 Iridium	Pt 78 Platinum	Au 79 Gold	Hg 80 Mercury	Tl 81 Thallium	Pb 82 Lead	Bi 83 Bismuth	Po 84 Polonium	At 85 Astatine	Rn 86 Radon
Fr 87 Francium	Ra 88 Radium		Rf 104 Rutherfordium	Db 105 Dubnium	Sg 106 Seaborgium	Bh 107 Bohrium	Hs 108 Hassium	Mt 109 Meitnerium	Ds 110 Darmstadtium	Rg 111 Roentgenium	Cn 112 Copernicium	Uut 113 Ununtrium	Fl 114 Flerovium	Uup 115 Ununpentium	Lv 116 Livermorium	Uus 117 Ununseptium	Uuo 118 Ununoctium
		<div> <div> Lanthanum 57 La </div> <div> Cerium 58 Ce </div> <div> Praseodymium 59 Pr </div> <div> Neodymium 60 Nd </div> <div> Promethium 61 Pm </div> <div> Samarium 62 Sm </div> <div> Europium 63 Eu </div> <div> Gadolinium 64 Gd </div> <div> Terbium 65 Tb </div> <div> Dysprosium 66 Dy </div> <div> Holmium 67 Ho </div> <div> Erbium 68 Er </div> <div> Thulium 69 Tm </div> <div> Ytterbium 70 Yb </div> <div> Lutetium 71 Lu </div> </div>															
		<div> <div> Actinium 89 Ac </div> <div> Thorium 90 Th </div> <div> Protactinium 91 Pa </div> <div> Uranium 92 U </div> <div> Neptunium 93 Np </div> <div> Plutonium 94 Pu </div> <div> Americium 95 Am </div> <div> Curium 96 Cm </div> <div> Berkelium 97 Bk </div> <div> Californium 98 Cf </div> <div> Einsteinium 99 Es </div> <div> Fermium 100 Fm </div> <div> Mendelevium 101 Md </div> <div> Nobelium 102 No </div> <div> Lawrencium 103 Lr </div> </div>															

Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

Section 2 nuclear Chemistry Study Guide Key

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