

\_\_\_\_\_

Identify the administrative variables that influence the business of managing the practice. (250-3 word summary paragraph per variable)

8. 10. How many moles of  $\text{H}_2\text{SO}_4$  are present in a solution of 100 g of concentrated  $\text{H}_2\text{SO}_4$ ?  
 (a) 1  
 (b) 1.1  
 (c) 1.2  
 (d) 1.3

9. 10. For concentrated  $\text{H}_2\text{SO}_4$ ,  $\rho = 1.84 \text{ g/cm}^3$ , what factor expresses its molarity?  
 (a) 18.4  
 (b) 18.3  
 (c) 18.2  
 (d) 18.1

10. 12. The equivalent weight of  $\text{H}_2\text{SO}_4$  is:  
 (a) half of its molecular weight  
 (b) equal to its molecular weight  
 (c) twice its molecular weight  
 (d) four times its molecular weight

11. 10. First ionization of  $\text{H}_2\text{SO}_4$  is 100% and second is 10%. How many moles of  $\text{H}^+$  ions are produced per mole of  $\text{H}_2\text{SO}_4$  in 1 M solution?  
 (a) 1.10 mole  
 (b) 1.05 mole  
 (c) 1.00 mole  
 (d) 0.95 mole

12. 10. What is the pH of the solution containing 0.01 mole of  $\text{H}_2\text{SO}_4$  in 100 ml of solution?  
 (a) 2.00  
 (b) 1.30  
 (c) 1.00  
 (d) 0.30

13. 10. What is the pH of the solution containing 0.01 mole of  $\text{H}_2\text{SO}_4$  in 100 ml of solution?  
 (a) 1.00  
 (b) 1.30  
 (c) 1.60  
 (d) 1.90

14. 10. What is the pH of the solution containing 0.01 mole of  $\text{H}_2\text{SO}_4$  in 100 ml of solution?  
 (a) 1.00  
 (b) 1.30  
 (c) 1.60  
 (d) 1.90

15. 10. A solution of  $\text{H}_2\text{SO}_4$  is 10% by weight. What is the molarity of the solution?  
 (a) 1.84  
 (b) 1.83  
 (c) 1.82  
 (d) 1.81

16. 10. First ionization of  $\text{H}_2\text{SO}_4$  is 100% and second is 10%. How many moles of  $\text{H}^+$  ions are produced per mole of  $\text{H}_2\text{SO}_4$  in 1 M solution?  
 (a) 1.10 mole  
 (b) 1.05 mole  
 (c) 1.00 mole  
 (d) 0.95 mole

17. 10. For concentrated  $\text{H}_2\text{SO}_4$ ,  $\rho = 1.84 \text{ g/cm}^3$ , what factor expresses its molarity?  
 (a) 18.4  
 (b) 18.3  
 (c) 18.2  
 (d) 18.1

18. 10. The equivalent weight of  $\text{H}_2\text{SO}_4$  is:  
 (a) half of its molecular weight  
 (b) equal to its molecular weight  
 (c) twice its molecular weight  
 (d) four times its molecular weight

19. 10. First ionization of  $\text{H}_2\text{SO}_4$  is 100% and second is 10%. How many moles of  $\text{H}^+$  ions are produced per mole of  $\text{H}_2\text{SO}_4$  in 1 M solution?  
 (a) 1.10 mole  
 (b) 1.05 mole  
 (c) 1.00 mole  
 (d) 0.95 mole

20. 10. What is the pH of the solution containing 0.01 mole of  $\text{H}_2\text{SO}_4$  in 100 ml of solution?  
 (a) 2.00  
 (b) 1.30  
 (c) 1.00  
 (d) 0.30

# Multiple Choice Stoichiometry Test

**Mary Millhollon, Richard H. Langley**



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