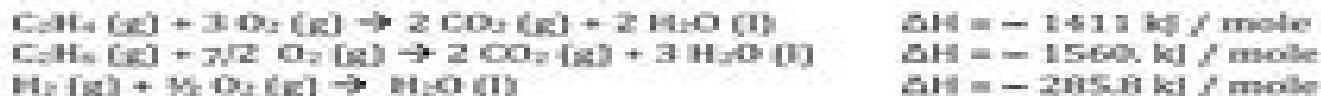


CHEM 12: Hess's Law Worksheet

1. Calculate the ΔH for the reaction: $\text{C}_2\text{H}_6(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g})$, from the following data:



2. Calculate the ΔH for the reaction:



3. Find the ΔH_f° (heat of formation) for acetic acid, $\text{HC}_2\text{H}_3\text{O}_2$, using the following thermochemical data:



Pure Factor $\text{C}(\text{s})$ is in the form of graphite.

Acetic acid, $\text{HC}_2\text{H}_3\text{O}_2$, actually has the structure: CH_3COOH



4. Calculate the ΔH for the reaction: $\text{CH}_4(\text{g}) + \text{NH}_3(\text{g}) \rightarrow \text{HCN}(\text{g}) + 3\text{H}_2(\text{g})$, from the following data:



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B Lingard

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