



# HYDROGEN IONS AND ACIDITY

## Section Summary

### Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

### Key Vocabulary

<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>	<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>	<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>
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### Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

### Key Figures

- Figure 1: pH scale
- Figure 2: Ionization of water

### Principles of Ionization

Acid dissociation is the process of ionization in which an acid dissociates into ions. The dissociation of an acid is represented by the following equation:

$\text{HA} \rightleftharpoons \text{H}^+ + \text{A}^-$  (acid dissociation)

The equilibrium constant for this dissociation is the acid dissociation constant,  $K_a$ . The equilibrium constant for the dissociation of a weak acid is given by the following equation:

$K_a = \frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]}$  (acid dissociation constant)

The  $K_a$  of a weak acid is based on the relative number of  $\text{H}^+$  ions and  $\text{A}^-$  ions in the solution.

$K_a$  is a measure of the strength of an acid.

### Acid Strength

$K_a > 10^{-1}$	Strong acid
$10^{-1} > K_a > 10^{-5}$	Medium acid
$10^{-5} > K_a > 10^{-10}$	Weak acid
$10^{-10} > K_a > 10^{-15}$	Very weak acid
$10^{-15} > K_a$	Nonacid

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**Ernest Henry Starling,Hamilton  
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