



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Define acid and base and relate them to the pH of a solution.
- Define strong and weak acids.

Key Vocabulary

<ul style="list-style-type: none">hydrogen ionacidbasestrong acidweak acidpH	<ul style="list-style-type: none">hydrogen ion concentrationacidicbasicneutraltitrationtitration curve
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Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

Key Diagrams

- Titration curve
- Titration curve

Points of Consideration

Acidic and basic properties are determined by the concentration of hydrogen ions in a solution. The concentration of hydrogen ions is measured by the pH of a solution.

Hydrogen ions are the only ions that contribute to the concentration of ions in a solution. The concentration of ions in a solution is determined by the concentration of hydrogen ions.

Hydrogen ions are the only ions that contribute to the concentration of ions in a solution. The concentration of ions in a solution is determined by the concentration of hydrogen ions.

Hydrogen ions are the only ions that contribute to the concentration of ions in a solution. The concentration of ions in a solution is determined by the concentration of hydrogen ions.

Key Equations

$\text{pH} = -\log_{10}[\text{H}^+]$	$\text{pH} = \log_{10}[\text{OH}^-]$
$\text{pH} = 7$	$\text{pH} = 14$
$\text{pH} = 1$	$\text{pH} = 13$
$\text{pH} = 2$	$\text{pH} = 12$
$\text{pH} = 3$	$\text{pH} = 11$
$\text{pH} = 4$	$\text{pH} = 10$
$\text{pH} = 5$	$\text{pH} = 9$
$\text{pH} = 6$	$\text{pH} = 8$
$\text{pH} = 7$	$\text{pH} = 7$
$\text{pH} = 8$	$\text{pH} = 6$
$\text{pH} = 9$	$\text{pH} = 5$
$\text{pH} = 10$	$\text{pH} = 4$
$\text{pH} = 11$	$\text{pH} = 3$
$\text{pH} = 12$	$\text{pH} = 2$
$\text{pH} = 13$	$\text{pH} = 1$
$\text{pH} = 14$	$\text{pH} = 0$

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web calculate the mass percent mole fraction molality and molarity of hexane c 6 h 14 d 0 660 g cm 3 when it is mixed with heptane c 7 h 12 d 0 684 g cm3 in a 35 0 55 0 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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web converting moles and mass the molecular weight of sodium chloride nacl is 58 44 g mol how many moles of salt are in 13 8 g of sodium chloride express the answer using 3 significant figures stuck use a hint learn for free about math art computer programming economics physics chemistry biology medicine finance history and more

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web a derivatization that achieves this goes as molarity moles of solute volume of solution in liter moles molarity x volume in liter mass moles x molar mass numerical answers 1

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web what is the mole fraction of licl in the solution 11 8 practice problem acetone c 3 h 6 o has a density of 0 791 g ml and xylene c 8 h 10 has a density of 0 864 g ml a solution was prepared by mixing acetone and xylene calculate the mole fraction of acetone in a solution if 19 0 ml of acetone was dissolved in 98 4 ml of xylene 7 1

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web since the mole fraction of the o 2 is 0 050 it contributes 0 050 or if you will 5 0 of the total pressure thus total pressure is 0 21 0 050 4 2 atm problem 9 a sample of 1 43 g of helium and an unweighed quantity of o 2 are mixed in a flask at room temperature