

# HYDROGEN IONS AND ACIDITY

## Section Summary

### Objectives

- Define hydrogen ion and acid.
- Explain how acids ionize in aqueous solutions.
- Describe strong and weak acids.

### Key Vocabulary

• acid	• hydrogen ion	• strong acid
• dissociation	• ionization	• weak acid
• aqueous solution	• pH scale	• pH indicator

### Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

### Principles Observed

Acid dissociation is the process of ionizing an acid in an aqueous solution. An acid dissociates when it is dissolved in water.

Water dissociates into a hydrogen ion and a hydroxide ion. The concentration of these ions is equal to one another. This is called a neutral solution.

Acids dissociate into hydrogen ions and anions. The concentration of hydrogen ions is greater than the concentration of the anions. This is called an acidic solution.

Base dissociation is the process of ionizing a base in an aqueous solution. The concentration of the hydroxide ions is greater than the concentration of the anions. This is called a basic solution.

Acids dissociate into hydrogen ions and anions.

Acid	Concentration of H <sup>+</sup>	Concentration of OH <sup>-</sup>	Concentration of Anions	Concentration of OH <sup>-</sup> > H <sup>+</sup>	Concentration of H <sup>+</sup> > OH <sup>-</sup>	Concentration of H <sup>+</sup> = OH <sup>-</sup>
strong acid	high	low	low	no	yes	no
weak acid	moderate	moderate	moderate	no	yes	no
base	low	high	high	yes	no	no
water	moderate	moderate	moderate	no	no	yes

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**Emilie Sanchez**

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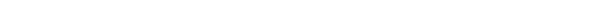
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